

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from pH and pOH , and vice versa from pH and pOH .
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. pH | 6. pOH |

Key Equation

- | | |
|------------------------------------|--------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pOH} = -\log[\text{OH}^-]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} + \text{pOH} = 14$ |

Part A Completion

Write the numerical value or chemical formula, of the missing quantity that is indicated in the space provided. Be sure to include proper units when appropriate.

1. What substance is _____, which hydrogen ion concentration is 1.0×10^{-7} M? The concentration of this acidic substance is 1.0×10^{-7} M.
2. Hydrochloric acid has a pH of _____, or a 1.0×10^{-2} M concentration. Calculate the pOH of this acid. The pOH is _____.
3. _____, which has a pH of _____, is a weak acid. The pOH of this acid is _____.
4. _____, which has a pH of _____, is a weak acid. The pOH of this acid is _____.
5. _____, which has a pH of _____, is a weak acid. The pOH of this acid is _____.
6. _____, which has a pH of _____, is a weak acid. The pOH of this acid is _____.
7. _____, which has a pH of _____, is a weak acid. The pOH of this acid is _____.

8. _____
9. _____
10. _____
11. _____
12. _____
13. _____
14. _____
15. _____
16. _____
17. _____
18. _____

19 2 Hydrogen Ions And Acidity Work Sheet Answers

W. M. Clark



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