

Limiting Reagent & Percent Yield Practice Worksheet

1. When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

a. Write the balanced equation for the reaction given above:



b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.362 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol CuCl}_2} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13 \text{ g NaCl}$$

$$20 \text{ g NaNO}_3 \times \frac{1 \text{ mol NaNO}_3}{84.997 \text{ g NaNO}_3} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol NaNO}_3} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13.8 \text{ g NaCl}$$

Only 13 g can be formed

c. What is the limiting reagent for the reaction in #2? **copper (II) chloride**

d. How many grams of copper (II) nitrate is formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.482 \text{ g CuCl}_2} \times \frac{1 \text{ mol Cu}(\text{NO}_3)_2}{1 \text{ mol CuCl}_2} \times \frac{187.467 \text{ g Cu}(\text{NO}_3)_2}{1 \text{ mol Cu}(\text{NO}_3)_2} = 20.9 \text{ g Cu}(\text{NO}_3)_2$$

e. How much of the excess reagent is left over in this reaction?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.482 \text{ g CuCl}_2} \times 2 \text{ mol NaNO}_3 \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 19 \text{ g NaNO}_3$$

$$20 \text{ g NaNO}_3 - 19 \text{ g NaNO}_3 = 1 \text{ g NaNO}_3$$

f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$$

$$11.3/13.0 \times 100\% = 86.9\%$$

2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.

a. Balance the following equation:



b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

$$25.0 \text{ g Pb}(\text{NO}_3)_2 \times \frac{1 \text{ mol Pb}(\text{NO}_3)_2}{331.214 \text{ g Pb}(\text{NO}_3)_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol Pb}(\text{NO}_3)_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 12.8 \text{ g NaNO}_3$$

$$15.0 \text{ g NaI} \times \frac{1 \text{ mol NaI}}{149.89 \text{ g NaI}} \times \frac{2 \text{ mol NaNO}_3}{2 \text{ mol NaI}} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

8.51 g NaNO₃ can be made

Answers Limiting Reactants And Percent Yield

Tracy Lynn Haroff

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introduce students to the equipment record keeping and safety of the laboratory Chapters 3 6 and 8 are designed to introduce students to laboratory techniques needed to perform all experiments In Chapters 7 and 9 through 20 students are required to use the techniques to synthesize compounds and analyze their properties In Chapter 21 students are introduced to multi step syntheses of organic compounds a practice well known in chemical industry In Chapter 23 students are asked to solve structures of unknown compounds The new chapter 24 introduces a meaningful experiment into the textbook that reflects the increasing emphasis on bioorganic chemistry in the sophomore level organic lecture course This experiment not only gives students the opportunity to accomplish a mechanistically interesting and synthetically important coupling of two a amino acids to produce a dipeptide but also provides valuable experience regarding the role of protecting groups in effecting synthetic transformations with multiple functionalized molecules *Let's Review* Albert S. Tarendash,1998 **5 Steps to a**

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1. Understanding the eBook Answers Limiting Reactants And Percent Yield
 - The Rise of Digital Reading Answers Limiting Reactants And Percent Yield
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Limiting Reactants And Percent Yield
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Limiting Reactants And Percent Yield
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Limiting Reactants And Percent Yield
 - Personalized Recommendations
 - Answers Limiting Reactants And Percent Yield User Reviews and Ratings
 - Answers Limiting Reactants And Percent Yield and Bestseller Lists
5. Accessing Answers Limiting Reactants And Percent Yield Free and Paid eBooks
 - Answers Limiting Reactants And Percent Yield Public Domain eBooks
 - Answers Limiting Reactants And Percent Yield eBook Subscription Services
 - Answers Limiting Reactants And Percent Yield Budget-Friendly Options
6. Navigating Answers Limiting Reactants And Percent Yield eBook Formats

- ePub, PDF, MOBI, and More
- Answers Limiting Reactants And Percent Yield Compatibility with Devices
- Answers Limiting Reactants And Percent Yield Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Answers Limiting Reactants And Percent Yield
- Highlighting and Note-Taking Answers Limiting Reactants And Percent Yield
- Interactive Elements Answers Limiting Reactants And Percent Yield

8. Staying Engaged with Answers Limiting Reactants And Percent Yield

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Answers Limiting Reactants And Percent Yield

9. Balancing eBooks and Physical Books Answers Limiting Reactants And Percent Yield

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Answers Limiting Reactants And Percent Yield

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Answers Limiting Reactants And Percent Yield

- Setting Reading Goals Answers Limiting Reactants And Percent Yield
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Answers Limiting Reactants And Percent Yield

- Fact-Checking eBook Content of Answers Limiting Reactants And Percent Yield
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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