

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



Reactant side
4 Fe atoms
4 S atoms
 $7(2) = 14$ O atoms

Product side
4 Fe atoms
4 S atoms
 $2(3) + 4(2) = 14$ O atoms

b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

$m \text{ FeS} = 2.0 \text{ g}$
Req'd. = $m \text{ O}_2$

Solution:

molar mass of $\text{FeS} = 88 \text{ g/mol}$
mole ratio = 4 mol FeS : 7 mol O_2
molar mass of $\text{O}_2 = 32 \text{ g/mol}$

$$m_{\text{O}_2} = 2.0 \text{ g FeS} \left(\frac{1 \text{ mol FeS}}{88 \text{ g FeS}} \right) \left(\frac{7 \text{ mol O}_2}{4 \text{ mol FeS}} \right) \left(\frac{32 \text{ g O}_2}{1 \text{ mol O}_2} \right) = 1.3 \text{ g O}_2$$

2. Rusting involves the reaction of iron, Fe with oxygen, O_2 to form iron (III) oxide, Fe_2O_3 . (20pts)

a) Write the balanced chemical equation.



Reactant side
4 Fe atoms
 $3(2) = 6$ O atoms

Product side
4 Fe atoms
 $2(3) = 6$ O atoms

b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

$$4 \text{ mol Fe} : 3 \text{ mol O}_2$$

Product to each reactant:

$$2 \text{ mol Fe}_2\text{O}_3 : 4 \text{ mol Fe}$$

Answer Key To Stoichiometry Homework Problems

David Wilson

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