

# 1999 Advanced Placement Chemistry Exam

## Section I: Multiple Choice

Time — 1 hour and 30 minutes

NO CALCULATORS MAY BE USED WITH SECTION I.

Note: For all questions, assume that the temperature is 298 K, the pressure is 1.00 atmosphere, and solutions are aqueous unless otherwise specified.

Throughout the test the following symbols have the definitions specified unless otherwise noted.

$T$  = temperature  
 $P$  = pressure  
 $V$  = volume  
 $S$  = entropy  
 $H$  = enthalpy  
 $G$  = free energy  
 $R$  = molar gas constant  
 $n$  = number of moles  
 $M$  = molar

$m$  = molal  
 $L, mL$  = liter(s), milliliter(s)  
 $g$  = gram(s)  
 $nm$  = nanometer(s)  
 $atm$  = atmosphere(s)  
 $J, kJ$  = joule(s), kilojoule(s)  
 $V$  = volt(s)  
 $mol$  = mole(s)

### Part A

Directions: Each set of lettered choices below refers to the numbered statements immediately following it. Select the one lettered choice that best fits each statement and then fill in the corresponding oval on the answer sheet. A choice may be used once, more than once, or not at all in each set.

Questions 1–4 refer to the following types of energy.

- (A) Activation energy
- (B) Free energy
- (C) Ionization energy
- (D) Kinetic energy
- (E) Lattice energy

1. The energy required to convert a ground-state atom in the gas phase to a gaseous positive ion
2. The energy change that occurs in the conversion of an ionic solid to widely separated gaseous ions
3. The energy in a chemical or physical change that is available to do useful work
4. The energy required to form the transition state in a chemical reaction

Questions 5–8 refer to atoms for which the occupied atomic orbitals are shown below.

- (A)  $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}}$
- (B)  $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}}$
- (C)  $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}} 2p \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$
- (D)  $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}} 2p \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$
- (E)  $[Ar] 4s \underline{\hspace{1cm}} 3d \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$

5. Represents an atom that is chemically unreactive
6. Represents an atom in an excited state
7. Represents an atom that has four valence electrons
8. Represents an atom of a transition metal

Questions 9–12 refer to aqueous solutions containing 1:1 mole ratios of the following pairs of substances. Assume all concentrations are 1  $M$ .

# Ap Chemistry Exam Multiple Choice Explanations

**Angela Woodward Spangenberg**



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