



According to the equation above, how many moles of potassium chlorate,  $\text{KClO}_3$ , must be decomposed to generate 1.0 L of  $\text{O}_2$  gas at standard temperature and pressure?

A)  $\frac{1}{3} \left( \frac{1}{22.4} \right)$  mol  
 B)  $\frac{1}{2} \left( \frac{1}{22.4} \right)$  mol  
 C)  $\frac{2}{3} \left( \frac{1}{22.4} \right)$  mol  
 D)  $\frac{3}{2} \left( \frac{1}{22.4} \right)$  mol  
 E)  $2 \left( \frac{1}{22.4} \right)$  mol

19. Which of the following ions has the same number of electrons as  $\text{Br}^-$ ?

A)  $\text{Cl}^2-$   
 B)  $\text{K}^+$   
 C)  $\text{Sr}^{2+}$   
 D)  $\text{F}^-$   
 E)  $\text{Cl}^-$

20. In solid methanes, the forces between neighboring  $\text{CH}_4$  molecules are best characterized as

A) ionic bonds  
 B) covalent bonds  
 C) hydrogen bonds  
 D) ion-dipole forces  
 E) London (dispersion) forces

21. Of the following electron configurations of neutral atoms, which represents an atom in an excited state?

A)  $1s^2 2s^2 2p^3$   
 B)  $1s^2 2s^2 2p^2 3s^1$   
 C)  $1s^2 2s^2 2p^6 3s^1$   
 D)  $1s^2 2s^2 2p^6 3s^1 3p^2$   
 E)  $1s^2 2s^2 2p^6 3s^1 3p^3$

22. Which of the following is a nonpolar molecule that contains polar bonds?

A)  $\text{P}_2$   
 B)  $\text{CHF}_3$   
 C)  $\text{CO}_2$   
 D)  $\text{HCl}$   
 E)  $\text{NH}_3$

23. The oxidation state that is common to aqueous ions of Fe, Mn, and Zn is

A) +1  
 B) +2  
 C) +3  
 D) +4  
 E) +5

24. Which of the following shows the correct number of protons, neutrons, and electrons in a neutral carbon-134 atom?

	Protons	Neutrons	Electrons
A)	55	55	55
B)	55	79	55
C)	55	79	79
D)	79	55	79
E)	134	55	134

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**Mary Millhollon, Richard H. Langley**

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