

7&8 • Atomic Structure & Periodicity

PRACTICE TEST

Rhe = A

$$\lambda = 2.18 \times 10^{-8} \text{ m}$$

$$R = 1.097 \times 10^7 \text{ m}^{-1}$$

$$\text{mass of an electron} = 9.11 \times 10^{-31} \text{ kg}$$

$$h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s}$$

$$c = 3.00 \times 10^8 \text{ m} \cdot \text{s}^{-1}$$

1. What wavelength corresponds to a frequency of $8.22 \times 10^9 \text{ Hz}$?
 a) 0.307 m d) 0.110 m
 b) 0.0365 m $\frac{3.00 \times 10^8 \text{ m}}{8.22 \times 10^9 \text{ s}^{-1}} = 27.4 \text{ m}$
 c) 0.122 m $\frac{3.00 \times 10^8 \text{ m}}{8.22 \times 10^9 \text{ s}^{-1}} = 0.36496 \text{ m}$

2. A radio station transmits at 110 MHz ($110 \times 10^6 \text{ Hz}$). What wavelength is this radio wave?
 a) $3.65 \times 10^{-5} \text{ m}$ c) $3.81 \times 10^{-5} \text{ m}$
 b) 3.30 m d) 2.73 m
 $\lambda = \frac{c}{\nu} = \frac{3.00 \times 10^8 \text{ m}}{110 \times 10^6 \text{ s}^{-1}} = 2.73 \text{ m}$

3. Which one of the following is NOT a proper unit for frequency?
 a) Hz c) $\text{m} \cdot \text{s}^{-1}$ *velocity*
 b) s^{-1} d) $\frac{1}{\text{sec}}$

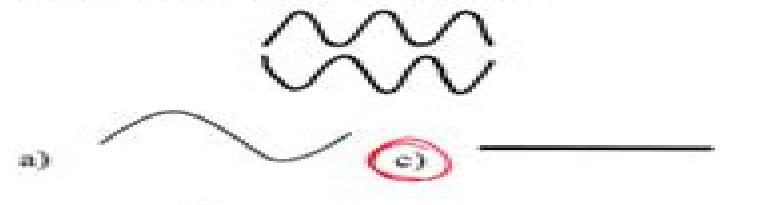
4. Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen spectrum.
 $E_4 = -2.18 \times 10^{-18} \text{ J}$ $E_2 = -2.18 \times 10^{-13} \text{ J}$
 $\Delta E = 4.84 \times 10^{-19} \text{ J}$ $\lambda = \frac{hc}{\Delta E} = \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s} \cdot 3.00 \times 10^8 \text{ m}}{4.84 \times 10^{-19} \text{ J}} = 4.1029 \times 10^{-7} \text{ m}$

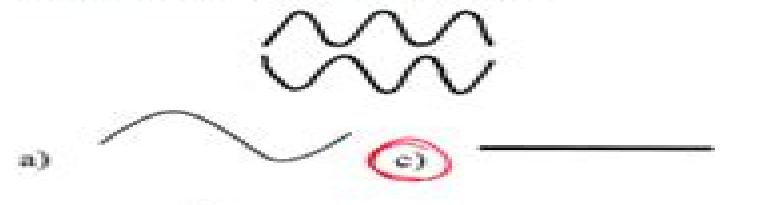
5. What is the relationship between the energy of a photon of light and its frequency?
 a) $E = \nu$ d) $E = \frac{1}{\nu}$
 b) $E = \frac{h}{\nu}$ c) $E = \frac{\nu}{h}$
 e) $E = h\nu$

6. What is the energy needed to raise an electron in the hydrogen atom from the second energy level to the third energy level?
 a) $1.52 \times 10^{-4} \text{ J}$ d) $4.48 \times 10^{-19} \text{ J}$
 b) $3.63 \times 10^{-19} \text{ J}$ e) $3.03 \times 10^{-19} \text{ J}$
 $\Delta E = E_3 - E_2 = \frac{hc}{\lambda_3} - \frac{hc}{\lambda_2} = \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s} \cdot 3.00 \times 10^8 \text{ m}}{9.14 \times 10^{-8} \text{ m}} - \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s} \cdot 3.00 \times 10^8 \text{ m}}{1.097 \times 10^7 \text{ m}} = 3.03 \times 10^{-19} \text{ J}$

7. What is the de Broglie wavelength of an electron moving at 80.0% the speed of light?
 a) $3.03 \times 10^{-12} \text{ m}$ c) $3.30 \times 10^{11} \text{ m}$
 b) $2.42 \times 10^{-12} \text{ m}$ d) $1.59 \times 10^{-25} \text{ m}$
 $\lambda = \frac{h}{mv} = \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s}}{7.14 \times 10^5 \text{ kg} \cdot 8.0 \times 3.00 \times 10^8 \text{ m/s}} = 8.8 \times 10^{-31} \text{ m}$

8. What resultant is expected from the interference of the two waves shown below?



a)  c) 
 b) d) 

9. Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?
 a) m_s c) m_l
 b) ℓ d) m_s

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1. Understanding the eBook Ap Chemistry Answer Key
 - The Rise of Digital Reading Ap Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ap Chemistry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chemistry Answer Key
 - Personalized Recommendations
 - Ap Chemistry Answer Key User Reviews and Ratings

- Ap Chemistry Answer Key and Bestseller Lists
- 5. Accessing Ap Chemistry Answer Key Free and Paid eBooks
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 - Participating in Virtual Book Clubs
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ap Chemistry Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ap Chemistry Answer Key
 - Setting Reading Goals Ap Chemistry Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ap Chemistry Answer Key
 - Fact-Checking eBook Content of Ap Chemistry Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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