

### Chemical Kinetics

1). Molecular hydrogen can be made from methane gas by the reaction below. How is the rate of disappearance of  $\text{CH}_4$  related to the rate of appearance of  $\text{H}_2$ ?



(A)   
 (B)   
 (C)   
 (D) none of these

Answer: B

2). For the reaction shown below, what is the relationship between the rate of formation of  $\text{Br}_2$  and the rate of reaction of  $\text{Br}^-$  over the same time period?



(A) rate of formation of  $\text{Br}_2 = 1/2 \times$  rate of reaction of  $\text{Br}^-$   
 (B) rate of formation of  $\text{Br}_2 = 2/4 =$  rate of reaction of  $\text{Br}^-$   
 (C) rate of formation of  $\text{Br}_2 = 2 \times$  rate of reaction of  $\text{Br}^-$   
 (D) rate of formation of  $\text{Br}_2 = 2 \times$  rate of reaction of  $\text{Br}^-$

Answer: A

3). The reaction that occurs in a Breathalyzer, a device used to determine the alcohol level in a person's bloodstream, is given below. If the rate of appearance of  $\text{Cr}_2(\text{SO}_4)_3$  is 3.20 mol/min at a particular moment, what is the rate of disappearance of  $\text{C}_2\text{H}_5\text{OH}$  at that moment?



(A) 0.400 mol/min  
 (B) 0.600 mol/min  
 (C) 1.66 mol/min  
 (D) 3.72 mol/min

Answer: C

# Ap Chem Multiple Choice

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