



REACTIONS OF ACIDS 1

metal + acid salt + hydrogen
 metal oxide + acid salt + water
 metal hydroxide + acid salt + water
 metal carbonate + acid salt + water + CO₂
 ammonia + acid ammonium salt

PART A Complete the table to show the name of the salt formed when the following acids react with the bases.

	nitric acid, HNO ₃	hydrochloric acid, HCl	sulfuric acid, H ₂ SO ₄
sodium carbonate, Na ₂ CO ₃	<u>sodium nitrate</u>	<u>sodium chloride</u>	<u>sodium sulfate</u>
magnesium, Mg	<u>magnesium nitrate</u>	<u>magnesium chloride</u>	<u>magnesium sulfate</u>
potassium oxide, K ₂ O	<u>potassium nitrate</u>	<u>potassium chloride</u>	<u>potassium sulfate</u>
copper hydroxide, Cu(OH) ₂	<u>copper nitrate</u>	<u>copper chloride</u>	<u>copper sulfate</u>
ammonia, NH ₃	<u>ammonium nitrate</u>	<u>ammonium chloride</u>	<u>ammonium sulfate</u>

PART B Now complete the following word equations for reactions between some acids and some bases.

- iron + hydrochloric acid iron chloride
- hydrochloric acid + copper carbonate copper chloride + water + carbon dioxide
- iron(II) hydroxide + sulfuric acid iron sulfate + water
- nitric acid + calcium oxide calcium nitrate + water
- sulfuric acid + ammonia ammonium sulfate
- zinc carbonate + sulfuric acid zinc sulfate + water + carbon dioxide
- magnesium + nitric acid magnesium nitrate + hydrogen
- ammonia + nitric acid ammonium nitrate
- hydrochloric acid + potassium oxide potassium chloride + _____
- calcium hydroxide + citric acid calcium citrate + _____

PART C On the back of the sheet, write a balanced equation for reactions 2-9.

Area	Strength	To develop	Area	Strength	To develop	Area	Strength	To develop
Done with care and thoroughness			word equations for acid + oxide			write formulae		
Good SPaG			word equations for acid + hydroxide			write balanced equations		
Can identify salts formed from acids			word equations for acid + carbonate					
word equations for acid + metal			word equations for acid + ammonia					

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