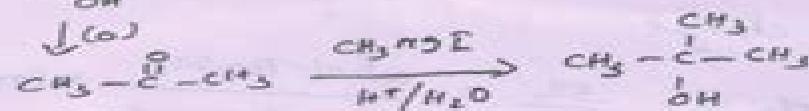


51) Pb^{+2} is more stable than Pb^{+4}

$FeCl_3$ changes to Pb^{+2} to Pb^{+4}

52) $P = P_{O_2} \times O_2 + P_{H_2} \times H_2$

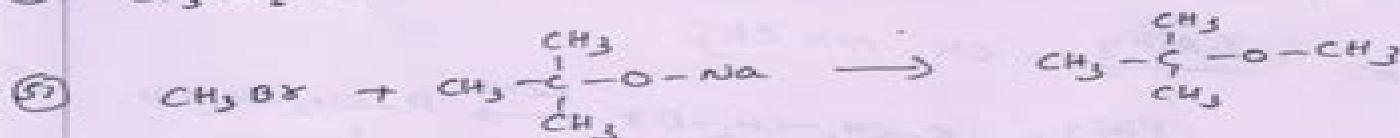
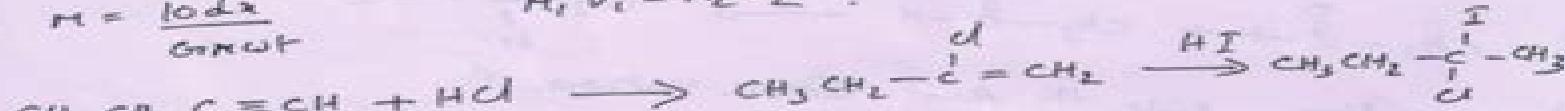
$$P = P_{O_2} + P_{H_2} \rightarrow P_{H_2} = \frac{742}{2} = 370$$



54) 1 SO^{+2} creates one cationic vacancy

$$\text{no of cationic vacancies} = \frac{10^{-4} \times 6 \times 10^{23}}{100} = 6 \times 10^{17}$$

55) $n = \frac{10dx}{G \cdot m \cdot t} \quad n_1, V_1 = n_2, V_2$



57) The non metal which R.P can displace over R.P non metal from their salt solution.



59) IA group carbonates more stable than IIA group carbonates.

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