

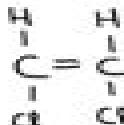
AP Chemistry

Session 5 – MCQ (Set-Based Questions)

Questions 1-3 refer to the reactions represented below, which are involved in a demonstration commonly known as “underwater fireworks.”



1. $\text{Ca}(\text{OH})_2$ precipitates when a 1.0 g sample of $\text{CaC}_2(s)$ is added to 1.0 L of distilled water at room temperature. If a 0.064 g sample of $\text{CaC}_2(s)$ (molar mass 64 g/mol) is used instead and all of it reacts, which of the following will occur and why? (The value of K_{sp} for is 8.0×10^{-6})
A. will precipitate because $Q > K_{\text{sp}}$
B. will precipitate because $Q < K_{\text{sp}}$
C. will not precipitate because $Q > K_{\text{sp}}$
 D. will not precipitate because $Q < K_{\text{sp}}$
2. Reaction 2 occurs when an excess of 6 M HCl(aq) solution is added to 100. mL of NaOCl(aq) of unknown concentration. If the reaction goes to completion and 0.010 mol of $\text{Cl}_2(g)$ is produced, then what was the molarity of the NaOCl(aq) solution?
A. 0.0010 M
B. 0.010 M
 C. 0.10 M
D. 1.0 M
3. When Reaction 3 occurs, does the hybridization of the carbon atoms change?
 A. Yes; it changes from sp to sp^2
B. Yes; it changes from sp to sp^3
C. Yes; it changes from sp^2 to sp^3
D. No; it does not change



20ap Chemistry Released Exam

Terry C. Jones

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