

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from pH and pOH , and vice versa from pH and pOH .
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. pH | 6. pOH |

Key Equation

- | | |
|-------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

Part A Completion

Write the hydrogen ion concentration or autoionization equilibrium of the acids and bases that are indicated in the spaces provided. Be sure to include the correct state symbols in the chemical equations.

1. What substance is _____, a weak hydrogen ion concentration of $1.0 \times 10^{-4} \text{ M}$ is the concentration of this acidic substance in a 0.10 M solution? _____
2. Hydrochloric acid ionizes to form _____ and _____.
3. _____, a weak acid, ionizes to form _____ and _____.
4. _____, a weak acid, ionizes to form _____ and _____.
5. _____, a weak acid, ionizes to form _____ and _____.
6. _____, a weak acid, ionizes to form _____ and _____.
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17. _____, a weak acid, ionizes to form _____ and _____.
18. _____, a weak acid, ionizes to form _____ and _____.
19. _____, a weak acid, ionizes to form _____ and _____.
20. _____, a weak acid, ionizes to form _____ and _____.

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19 2 Hydrogen Ions And Acidity Worksheet Answers

Yi-Tong Ma

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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