

**Analytical Chemistry**  
**3<sup>rd</sup> Long Exam**

**PART I. MULTIPLE CHOICE.** Choose the letter that corresponds to the BEST answer. Write E when the answer is NONE in the choices. (25 pts)

1. Transfer of electrons from one reactant to another

- a. precipitation rxn                      c. redox rxn  
b. half rxn                                  d. neutralization rxn

2. What is the mole ratio of  $\text{Ce(IV)}$  to  $\text{NO}_2^-$  for the oxidation of nitrite to nitrate using  $\text{Ce}^{4+}$  in the presence of an acid:



- a. 1:2    c. 1:1  
b. 2:1    d. 2:2

3. How many mole(s) of electron is(are) produced in the oxidation of sodium bisulfate ( $\text{NaHSO}_4$ ) to  $\text{BiO}^+$  and  $\text{H}^+$  under acidic conditions?

- a. 1    c. 3  
b. 2    d. 4

4. Azide ( $\text{N}_3^-$ ) was added in the determination of the dissolved oxygen in water sample to remove interferences caused by the presence of what ion?

- a. chloride                                  c. nitrite  
b. fluoride                                  d. nitrate

5. Which of the following statements is(are) true?

- I. Prepared  $\text{KMnO}_4$  solutions should be stored in amber bottles  
II.  $\text{KMnO}_4$  solutions are boiled to remove  $\text{CO}_2$  which makes the solution basic  
III.  $\text{Mn(II)}$  is the product of the degradation of  $\text{KMnO}_4$   
IV.  $\text{KMnO}_4$  solutions are filtered to remove manganese(IV) oxide

- a. I only                                      c. I and IV  
b. I and II                                  d. I, II and III

6. What is the oxidation state of carbon in oxalic acid?

- a. +2    c. +6  
b. +4    d. +7

7. The complex formed between  $\text{Cu(I)}$  and 1,10-phenanthroline has a molar absorptivity of  $7000 \text{ L cm}^{-1} \text{ mol}^{-1}$  at 435 nm, the wavelength of maximum absorption. What is the absorbance of a  $5.17 \times 10^{-3} \text{ M}$  solution of the complex when measured in a 1.00-cm cell at 435 nm?

- a. 0.0431                                  c. 0.432  
b. 0.0432                                  d. 0.431

8. In the presence of iodine, thiosulfate ion is oxidized quantitatively to tetrathionate ion. What is the chemical formula of tetrathionate ion?

- a.  $\text{S}_4\text{O}_7^{2-}$                                   c.  $\text{S}_4\text{O}_6^{2-}$   
b.  $\text{S}_4\text{O}_5^{2-}$                                   d.  $\text{S}_4\text{O}_8^{2-}$

9. In Winkler situation, the flocc is produced from

- a. oxidation of  $\text{I}^-$  to  $\text{I}_2$                       c. reduction of  $\text{I}^-$  to  $\text{I}_2$   
b. oxidation of  $\text{Mn(II)}$  to  $\text{Mn(III)}$               d. reduction of  $\text{Mn(II)}$  to  $\text{Mn(III)}$

10. A component of a UV-Vis spectrophotometer that restricts the radiation being measured to a narrow band that is absorbed or emitted by the analyte.

- a. light source                              c. radiation filter  
b. detector                                  d. wavelength selector

11. Component of a UV-Vis instrument that operates by blocking or absorbing all but a restricted band of radiation.

- a. light source                              c. radiation filter  
b. detector                                  d. wavelength selector

12. Oxidation of  $\text{VO}^{2+}$  using permanganate or cerium as standard solution produces  $\text{VO(OH)}_2^{+}$ , what is the oxidation state of vanadium in  $\text{VO(OH)}_2^{+}$ ?

# Analytical Chemistry Multiple Choice Exam Questions

**Ralph H. Petrucci, Robert K. Wismer**



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