Chem 115: Answers to Chapter 6_Aqueous Ionic Equilibrium Practice Questions

Name of Instructor: Pearson W. K. Ahiahonu

Practice Question 1

Calculate the pH of the buffer that results from mixing 60.0 mL of 0.250 M $HCHO_2$ and 15.0 mL of 0.500 M $NaCHO_2$ (K_a for $HCHO_2 = 1.8 \times 10^{-4}$)

$$HCHO_{2\,(aq)} + H_2O_{(l)} \rightleftharpoons H_3O^+_{(aq)} + CHO_2^-_{(aq)}$$

$$\text{Total volume of buffer} = 60.0\,mL + 15.0\,mL$$
$$= 75.0\,mL = 0.075\,L$$

$$[HCHO_2] = \frac{0.060\,L}{0.075\,L} \times 0.250\,M = 0.200\,M$$

$$[CHO_2^-] = \frac{0.015\,L}{0.075\,L} \times 0.500\,M = 0.100\,M$$

$$\Rightarrow pH = pKa + \log \frac{[CHO_2^-]}{[HCHO_2]}$$

$$= -\log(1.8 \times 10^{-4}) + \log\left(\frac{0.100}{0.200}\right)$$

$$= -(0.25527 - 4) + (-0.30103)$$

$$= 3.74 47275 - 0.30103$$

$$= 3.44$$

1

Aqueous Equilibrium Practice Problems

Jianjun Gao

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1. Understanding the eBook Aqueous Equilibrium Practice Problems
    - The Rise of Digital Reading Aqueous Equilibrium Practice Problems
    - Advantages of eBooks Over Traditional Books
2. Identifying Aqueous Equilibrium Practice Problems
    - Exploring Different Genres
    - Considering Fiction vs. Non-Fiction
    - Determining Your Reading Goals
3. Choosing the Right eBook Platform
    - Popular eBook Platforms
    - Features to Look for in an Aqueous Equilibrium Practice Problems
    - User-Friendly Interface
4. Exploring eBook Recommendations from Aqueous Equilibrium Practice Problems
    - Personalized Recommendations
    - Aqueous Equilibrium Practice Problems User Reviews and Ratings
    - Aqueous Equilibrium Practice Problems and Bestseller Lists
5. Accessing Aqueous Equilibrium Practice Problems Free and Paid eBooks

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