

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the K_a and K_b of acids and bases and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-----------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. acid rain | 6. acid rain solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Read the paragraph and use the information to determine if the sentence is true or false. If the sentence is false, correct the statement so it becomes a true statement. Write the corrected statement in the space provided.

1. When autoionization is 1.0×10^{-14} , the concentration of hydronium ions is 1.0×10^{-7} M. The concentration of hydroxide ions is 1.0×10^{-7} M.
2. The pH of a solution is the negative logarithm of the concentration of hydronium ions. The pH of a solution is 1.0×10^{-14} .
3. The pH of a solution is the negative logarithm of the concentration of hydroxide ions. The pH of a solution is 1.0×10^{-14} .
4. The pH of a solution is the negative logarithm of the concentration of hydronium ions. The pH of a solution is 1.0×10^{-14} .

5. The pH of a solution is the negative logarithm of the concentration of hydronium ions. The pH of a solution is 1.0×10^{-14} .
6. The pH of a solution is the negative logarithm of the concentration of hydronium ions. The pH of a solution is 1.0×10^{-14} .

7. The pH of a solution is the negative logarithm of the concentration of hydronium ions. The pH of a solution is 1.0×10^{-14} .
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12. The pH of a solution is the negative logarithm of the concentration of hydronium ions. The pH of a solution is 1.0×10^{-14} .
13. The pH of a solution is the negative logarithm of the concentration of hydronium ions. The pH of a solution is 1.0×10^{-14} .
14. The pH of a solution is the negative logarithm of the concentration of hydronium ions. The pH of a solution is 1.0×10^{-14} .

19 2 Hydrogen Ions And Acidity Worksheet Answers

**La Motte Chemical Products
Company, William Anthony Taylor**



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