

Name: ANSWERS

## Ionic Bonding Worksheet

For each pair of elements, use Lewis Dot Diagrams to show the transfer of electrons from the metal(s) to the non-metal(s) and draw the final Lewis Dot Diagram (with square brackets and ionic charges) representing the ionic compound. In addition, write the formula and name of the resulting compound.

Li + Cl	$\text{Li} \curvearrowright \ddot{\text{Cl}}: \rightarrow [\text{Li}]^{1+} [\ddot{\text{Cl}}:]^{1-}$	Formula and Name lithium chloride $\text{LiCl}$
Ca + O	$\text{Ca} \curvearrowright \ddot{\text{O}}: \rightarrow [\text{Ca}]^{2+} [\ddot{\text{O}}:]^{2-}$	Formula and Name calcium oxide $\text{CaO}$
Al + Cl	$\begin{array}{c} \ddot{\text{Cl}}: \\   \\ \text{Al} \curvearrowright \ddot{\text{Cl}}: \\   \\ \ddot{\text{Cl}}: \end{array} \rightarrow \begin{array}{c} [\ddot{\text{Cl}}:]^{-} \\   \\ [\text{Al}]^{3+} \\   \\ [\ddot{\text{Cl}}:]^{-} \end{array}$	Formula and Name aluminum chloride $\text{AlCl}_3$
Na + O	$\text{Na} \curvearrowright \ddot{\text{O}}: \rightarrow [\text{Na}]^{1+} [\ddot{\text{O}}:]^{2-} [\text{Na}]^{1+}$	Formula and Name sodium oxide $\text{Na}_2\text{O}$
Al + Br	$\begin{array}{c} \ddot{\text{Br}}: \\   \\ \text{Al} \curvearrowright \ddot{\text{Br}}: \\   \\ \ddot{\text{Br}}: \end{array} \rightarrow \begin{array}{c} [\ddot{\text{Br}}:]^{-} \\   \\ [\text{Al}]^{3+} \\   \\ [\ddot{\text{Br}}:]^{-} \end{array}$	Formula and Name aluminum bromide $\text{AlBr}_3$
Mg + S	$\text{Mg} \curvearrowright \ddot{\text{S}}: \rightarrow [\text{Mg}]^{2+} [\ddot{\text{S}}:]^{2-}$	Formula and Name magnesium sulphide $\text{MgS}$

## 6 Ionic Bonding Answers

**JL Elias**



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