

CHEMISTRY 3.5ANSWERSWORKSHEET FIVEORGANICHaloalkanes

- 2-chlorobutane
 - 2,3-dichlorobutane
- 2-bromo, 2-methylpropane (3°)
 - 2-bromopropane (2°)
 - 2-chlorobutane (2°)
 - 1-chlorobutane (1°)
 - 2-bromo, 2-methylbutane (3°)(f)
 - 2-bromohexane (2°)
- Haloalkanes have polar C-X bonds which causes them to have some attraction for polar water molecules. However they cannot hydrogen bond to water like alcohols, and so the water molecules prefer to bond to each other rather than to the haloalkane molecules. So the haloalkanes tend to remain in an undissolved layer in the presence of water.
- The boiling points of similar molecular mass organic compounds reflect the strengths of their intermolecular forces. Alkanes have only weak Van der Waals forces between molecules so they are easily separated and have low boiling points. Haloalkanes have polar molecules and so there are stronger permanent dipole interactions between molecules raising their boiling points. Alcohol molecules can hydrogen bond to each other and these strong intermolecular forces give them higher boiling points still.
- $\text{CH}_3\text{CH}_2\text{Cl} + \text{OH}^- \rightarrow \text{CH}_3\text{CH}_2\text{OH} + \text{Cl}^-$
 - If the haloalkane was reacted with aqueous ammonia, the OH^- ions present would substitute the halogen and form an alcohol instead of the desired amine. In alcohol, the ammonia molecules do not react to produce OH^- ions and so they substitute successfully to make the amine.
 - $\text{CH}_3\text{CH}_2\text{Cl} + \text{NH}_3 \rightarrow \text{CH}_3\text{CH}_2\text{NH}_2 + \text{HCl}$
- $\text{CH}_3\text{CH}_2\text{CH(OH)CH}_3$
 - $\text{CH}_3\text{CH=CHCH}_3$
 - $\text{CH}_3\text{CH}_2\text{CHClCH}_3$
 - Substitution
 - Elimination
- $\text{CH}_3\text{CCl(CH}_3)_2 \xrightarrow{\text{Aqueous KOH}} \text{CH}_3\text{COH(CH}_3)_2$
 - $\text{CH}_3\text{CHBrCH}_3 \xrightarrow{\text{Alcoholic KOH}} \text{CH}_3\text{CH=CH}_2$
 - $\text{CH}_3\text{CH}_2\text{CH}_2\text{Cl} \xrightarrow{\text{Alcoholic NH}_3} \text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2$
 - $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{Cl} \xrightarrow{\text{Aqueous NaOH}} \text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$
- Aqueous KOH
 - Alcoholic NH_3
 - Alcoholic KOH

30 N Answers

Rosina Ehmann



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Table of Contents 3o N Answers

1. Understanding the eBook 3o N Answers
 - The Rise of Digital Reading 3o N Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 3o N Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 3o N Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 3o N Answers
 - Personalized Recommendations

- 3o N Answers User Reviews and Ratings
- 3o N Answers and Bestseller Lists
- 5. Accessing 3o N Answers Free and Paid eBooks
 - 3o N Answers Public Domain eBooks
 - 3o N Answers eBook Subscription Services
 - 3o N Answers Budget-Friendly Options
- 6. Navigating 3o N Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 3o N Answers Compatibility with Devices
 - 3o N Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 3o N Answers
 - Highlighting and Note-Taking 3o N Answers
 - Interactive Elements 3o N Answers
- 8. Staying Engaged with 3o N Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 3o N Answers
- 9. Balancing eBooks and Physical Books 3o N Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 3o N Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 3o N Answers
 - Setting Reading Goals 3o N Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 3o N Answers
 - Fact-Checking eBook Content of 3o N Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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