

**17 • Acid-Base Equilibria**

1. What is the  $[H^+]$  when  $[OH^-] = 8.1 \times 10^{-3}$ ?

- $8.1 \times 10^{-3} M$
- $1.0 \times 10^{-7} M$
- $1.2 \times 10^{-20} M$   $(H^+) = \frac{K_w}{[OH^-]} = \frac{1.0 \times 10^{-14}}{8.1 \times 10^{-3}} =$
- $3.6 \times 10^{-6} M$
- $8.1 \times 10^{-3} M$

2. What is the  $[H^+]$  when  $[OH^-] = 3.3 \times 10^{-7}$ ?

- $3.0 \times 10^{-6} M$
- $1.0 \times 10^{-7} M$
- $3.3 \times 10^{-7} M$   $(H^+) = \frac{K_w}{[OH^-]} = \frac{1.0 \times 10^{-14}}{3.3 \times 10^{-7}} =$
- $6.6 \times 10^{-8} M$
- $3.3 \times 10^{-7} M$

3. What is the  $[H^+]$  in a  $0.0025 M$  HCl solution?

- $1.0 \times 10^{-3} M$
- $4.0 \times 10^{-12} M$
- $2.5 \times 10^{-3} M$   $(H^+) = \frac{K_w}{[OH^-]} = \frac{1.0 \times 10^{-14}}{0.0025} = 2.5 \times 10^{-12} M$
- $3.6 \times 10^{-5} M$

4. What is the  $[OH^-]$  in a  $0.0050 M$  HCl solution?

- $5.0 \times 10^{-3} M$
- $1.0 M$
- $1.0 \times 10^{-7} M$   $(H^+) = \frac{K_w}{[OH^-]} = \frac{1.0 \times 10^{-14}}{0.0050} =$
- $6.6 \times 10^{-9} M$

5. A solution in which  $[H^+] = 10^{-8}$  has a pH of B and is basic.

- 8, acidic
- 6, basic
- 6, basic
- 8, neutral
- 8, basic

6. What is the pH of a  $0.00030 M$   $HNO_3$  solution?

- 8.11
- 3.00
- 3.52  $(H^+) = 3.0 \times 10^{-4}$
- 4.48
- none of these

7. What is the pH of a  $0.0060 M$  KOH solution?

- 5.12
- 2.22
- 11.78  $(OH^-) = 6.0 \times 10^{-4}$
- 8.88
- 7.00

post = 2.22  
pH = 14 - pOH =

**PRACTICE TEST**

8. A sample of lemon juice is found to have a pH of 2.55. What is the  $H^+$  concentration of the juice?  $(H^+) = 10^{-14} = 10^{-2.55}$

- $0.0035 M$
- $0.0028 M$
- $355 M$
- $0.0080 M$
- $11.6 M$

9. A sample of milk is found to have a pH of 6.60. What is the  $OH^-$  concentration of the milk?

- $2.5 \times 10^{-8} M$
- $1.0 \times 10^{-7} M$
- $5.0 \times 10^{-7} M$   $(H^+) = 10^{-6.60} = 2.5 \times 10^{-7}$
- $4.0 \times 10^{-8} M$
- $2.5 \times 10^{-7} M$

10. What is the concentration of  $Cl^-$  in a  $0.60 M$  solution of  $HOCl$ ?  $K_a = 3.1 \times 10^{-8}$

- $1.8 \times 10^{-4} M$
- $7.1 \times 10^{-11} M$
- $0.40 M$   $(OH^-) = (H^+) = Ka \cdot x = x^2$
- $1.4 \times 10^{-4} M$
- $1.1 \times 10^{-4} M$

see scratch paper

11. What is the pH of a  $0.020 M$  solution of hydrochloric acid, a diprotic acid?  $K_{a1} = 1.1 \times 10^{-7}$   $K_{a2} = 1.0 \times 10^{-14}$   $\text{dissociation}$   $\text{is significant}$   $\text{use } K_{a1}$   $\text{only}$

- 7.00
- 9.67
- 3.05
- 4.33
- 7.84

see scratch paper

12. What is the concentration of  $CO_3^{2-}$  in a  $0.010 M$  solution of carbonic acid? The relevant equilibria are:

$$H_2CO_3 \rightleftharpoons H^+ + HCO_3^- \quad K_{a1} = 4.3 \times 10^{-7}$$

$$HCO_3^- \rightleftharpoons H^+ + CO_3^{2-} \quad K_{a2} = 5.6 \times 10^{-11}$$

- $6.6 \times 10^{-5} M$
- $5.6 \times 10^{-11} M$
- $6.7 \times 10^{-11} M$
- $7.5 \times 10^{-7} M$
- $7.9 \times 10^{-7} M$

from the 1st dissociation,  $[H^+] = [HCO_3^-]$   
the 2nd dissociation doesn't change this much.  
 $K_{a2} = \frac{[H^+][CO_3^{2-}]}{[HCO_3^-]}$  but  $[H^+] = [HCO_3^-]$   $(H_2CO_3)$  is small!

so...  $K_{a2} = [CO_3^{2-}] = 5.6 \times 10^{-11}$

# Answer Key To Acids Bases And Solutions

**John A. Olmsted, Gregory M.  
Williams, Robert Charles Burk**

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