

Calculating pH

How is pH related to the concentration of hydronium ions?

Why?

In biology and other science courses pH is introduced as a way to quantify the acidity or basicity of a solution. This property can be measured using a pH probe or with an indicator paper strip that changes color at different pH values. But, what is actually being measured? We know that a pH of 7 is neutral, below 7 is acid, and above 7 is base, but why? What in the solution is the paper strip or probe actually reacting with?

Model 1 – Ion Concentrations for Acids and Bases

Beaker	Solution	Acidic, Basic or Neutral?	$[\text{H}_3\text{O}^+]$	$[\text{OH}^-]$	$[\text{H}_3\text{O}^+] \times [\text{OH}^-]$
1	0.10 M HCl(aq)	Acidic	$1.0 \times 10^{-1} \text{ M}$	$1.0 \times 10^{-13} \text{ M}$	1×10^{-14}
2	0.0010 M HCl(aq)	Acidic	$1.0 \times 10^{-2} \text{ M}$	$1.0 \times 10^{-12} \text{ M}$	1×10^{-14}
3	0.000010 M HCl(aq)	Acidic	$1.0 \times 10^{-3} \text{ M}$	$1.0 \times 10^{-11} \text{ M}$	1×10^{-14}
4	0.0000010 M HCl(aq)	Acidic	$1.0 \times 10^{-4} \text{ M}$	$1.0 \times 10^{-10} \text{ M}$	1×10^{-14}
5	0.00000010 M HCl(aq)	Neutral	$1.0 \times 10^{-5} \text{ M}$	$1.0 \times 10^{-9} \text{ M}$	1×10^{-14}
6	0.00000010 M NaOH(aq)	Neutral	$1.0 \times 10^{-7} \text{ M}$	$1.0 \times 10^{-7} \text{ M}$	1×10^{-14}
7	0.0000010 M NaOH(aq)	Basic	$1.0 \times 10^{-8} \text{ M}$	$1.0 \times 10^{-6} \text{ M}$	1×10^{-14}
8	0.00010 M NaOH(aq)	Basic	$1.0 \times 10^{-9} \text{ M}$	$1.0 \times 10^{-5} \text{ M}$	1×10^{-14}
9	0.010 M NaOH(aq)	Basic	$1.0 \times 10^{-10} \text{ M}$	$1.0 \times 10^{-4} \text{ M}$	1×10^{-14}
10	0.10 M NaOH(aq)	Basic	$1.0 \times 10^{-11} \text{ M}$	$1.0 \times 10^{-3} \text{ M}$	1×10^{-14}

1. What does the symbol $[\text{H}_3\text{O}^+]$ in Model 1 indicate?

The molar concentration of hydronium ion.

2. In Beaker 2, which ion has a higher concentration, hydronium ion or hydroxide ion?

Hydronium ion.

3. Describe how the concentration of hydronium ions was calculated for Beaker 3 in Model 1 from the concentration of the acid.

Since HCl is a strong acid, all of the molecules ionize in water. Therefore, the H_3O^+ concentration is equal to the initial acid concentration in the solution.

4. Describe how the concentration of hydroxide ion was calculated for Beaker 8 in Model 1 from the concentration of the base.

Since NaOH is a strong base, it completely dissociates in water. Therefore, the OH^- concentration is equal to the initial base concentration in the solution.

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**Juliette Lantz, Renée Cole, The POGIL
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 - The Rise of Digital Reading Answer Key Pogil Chemistry Calculating Ph
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Pogil Chemistry Calculating Ph
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Pogil Chemistry Calculating Ph
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Pogil Chemistry Calculating Ph
 - Personalized Recommendations
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 - Answer Key Pogil Chemistry Calculating Ph and Bestseller Lists
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- Answer Key Pogil Chemistry Calculating Ph Budget-Friendly Options

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- Answer Key Pogil Chemistry Calculating Ph Enhanced eBook Features

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- Interactive Elements Answer Key Pogil Chemistry Calculating Ph

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- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Answer Key Pogil Chemistry Calculating Ph

9. Balancing eBooks and Physical Books Answer Key Pogil Chemistry Calculating Ph

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Answer Key Pogil Chemistry Calculating Ph

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Answer Key Pogil Chemistry Calculating Ph

- Setting Reading Goals Answer Key Pogil Chemistry Calculating Ph
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Answer Key Pogil Chemistry Calculating Ph

- Fact-Checking eBook Content of Answer Key Pogil Chemistry Calculating Ph
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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