

## Calculating pH

How is pH related to the concentration of hydronium ions?

### Why?

In biology and other science courses pH is introduced as a way to quantify the acidity or basicity of a solution. This property can be measured using a pH probe or with an indicator paper strip that changes color at different pH values. But, what is actually being measured? We know that a pH of 7 is neutral, below 7 is acid, and above 7 is base, but why? What in the solution is the paper strip or probe actually reacting with?

### Model 1 – Ion Concentrations for Acids and Bases

Beaker	Solution	Acidic, Basic or Neutral?	$[H_3O^+]$	$[OH^-]$	$[H_3O^+] \times [OH^-]$
1	0.10 M HCl(aq)	Acidic	$1.0 \times 10^{-1}$ M	$1.0 \times 10^{-13}$ M	$1 \times 10^{-14}$
2	0.0010 M HCl(aq)	Acidic	$1.0 \times 10^{-3}$ M	$1.0 \times 10^{-11}$ M	$1 \times 10^{-14}$
3	0.000010 M HCl(aq)	Acidic	$1.0 \times 10^{-5}$ M	$1.0 \times 10^{-9}$ M	$1 \times 10^{-14}$
4	0.0000010 M HCl(aq)	Acidic	$1.0 \times 10^{-6}$ M	$1.0 \times 10^{-8}$ M	$1 \times 10^{-14}$
5	0.00000010 M HCl(aq)	Neutral	$1.0 \times 10^{-7}$ M	$1.0 \times 10^{-7}$ M	$1 \times 10^{-14}$
6	0.00000010 M NaOH(aq)	Neutral	$1.0 \times 10^{-7}$ M	$1.0 \times 10^{-7}$ M	$1 \times 10^{-14}$
7	0.0000010 M NaOH(aq)	Basic	$1.0 \times 10^{-8}$ M	$1.0 \times 10^{-6}$ M	$1 \times 10^{-14}$
8	0.00010 M NaOH(aq)	Basic	$1.0 \times 10^{-10}$ M	$1.0 \times 10^{-4}$ M	$1 \times 10^{-14}$
9	0.010 M NaOH(aq)	Basic	$1.0 \times 10^{-12}$ M	$1.0 \times 10^{-2}$ M	$1 \times 10^{-14}$
10	0.10 M NaOH(aq)	Basic	$1.0 \times 10^{-13}$ M	$1.0 \times 10^{-1}$ M	$1 \times 10^{-14}$

1. What does the symbol  $[H_3O^+]$  in Model 1 indicate?

*The molar concentration of hydronium ion.*

2. In Beaker 2, which ion has a higher concentration, hydronium ion or hydroxide ion?

*Hydronium ion.*

3. Describe how the concentration of hydronium ion was calculated for Beaker 3 in Model 1 from the concentration of the acid.

*Since HCl is a strong acid, all of the molecules ionize in water. Therefore, the  $H_3O^+$  concentration is equal to the initial acid concentration in the solution.*

4. Describe how the concentration of hydroxide ion was calculated for Beaker 8 in Model 1 from the concentration of the base.

*Since NaOH is a strong base, it completely dissociates in water. Therefore, the  $OH^-$  concentration is equal to the initial base concentration in the solution.*

# Answer Key Pogil Chemistry Calculating Ph

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