

# ANSWER KEY - ORGANIZED BY REACTION TYPE

Practice Bb	Practice Bbb	Practice Balancing Chemical Equations
<b>Part A</b> Instructions: Balance each of the following chemical equations. Write the correct coefficients in front of the number of atoms for each element. Do not change subscripts. (Show your work.) <ol style="list-style-type: none"> <li><math>1. \text{C} + \text{O}_2 \rightarrow \text{CO}_2</math> Balanced Equation</li> <li><math>2. \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}</math> Balanced Equation</li> <li><math>3. \text{H}_2 + \text{Cl}_2 \rightarrow \text{HCl}</math> Balanced Equation</li> <li><math>4. \text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3</math> Balanced Equation</li> <li><math>5. \text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}</math> Balanced Equation</li> </ol>	<b>Part B</b> Instructions: Balance each of the following chemical equations. Write the correct coefficients in front of the number of atoms for each element. Do not change subscripts. (Show your work.) <ol style="list-style-type: none"> <li><math>1. \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}</math> Balanced Equation</li> <li><math>2. \text{H}_2 + \text{Cl}_2 \rightarrow \text{HCl}</math> Balanced Equation</li> <li><math>3. \text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3</math> Balanced Equation</li> <li><math>4. \text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3</math> Balanced Equation</li> <li><math>5. \text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}</math> Balanced Equation</li> </ol>	<b>Part C: Double Replacement Reactions</b> Instructions: Balance each of the following chemical equations. Write the correct coefficients in front of the number of atoms for each element. Do not change subscripts. (Show your work.) <ol style="list-style-type: none"> <li><math>1. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math> Balanced Equation</li> <li><math>2. \text{KCl} + \text{AgNO}_3 \rightarrow \text{KNO}_3 + \text{AgCl}</math> Balanced Equation</li> <li><math>3. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math> Balanced Equation</li> <li><math>4. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math> Balanced Equation</li> <li><math>5. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math> Balanced Equation</li> </ol>

Practice Bb	Practice Balancing Chemical Equations
<b>Part C: Single Replacement Reactions</b> Instructions: Balance each of the following chemical equations. Write the correct coefficients in front of the number of atoms for each element. Do not change subscripts. (Show your work.) <ol style="list-style-type: none"> <li><math>1. \text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2</math> Balanced Equation</li> <li><math>2. \text{Fe} + \text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2</math> Balanced Equation</li> <li><math>3. \text{Fe} + \text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2</math> Balanced Equation</li> <li><math>4. \text{Fe} + \text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2</math> Balanced Equation</li> <li><math>5. \text{Fe} + \text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2</math> Balanced Equation</li> </ol>	<b>Part D: Double Replacement Reactions</b> Instructions: Balance each of the following chemical equations. Write the correct coefficients in front of the number of atoms for each element. Do not change subscripts. (Show your work.) <ol style="list-style-type: none"> <li><math>1. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math> Balanced Equation</li> <li><math>2. \text{KCl} + \text{AgNO}_3 \rightarrow \text{KNO}_3 + \text{AgCl}</math> Balanced Equation</li> <li><math>3. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math> Balanced Equation</li> <li><math>4. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math> Balanced Equation</li> <li><math>5. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math> Balanced Equation</li> </ol>

## ANSWER KEY

### Part A: Combustion Reactions

- $1. \text{C} + \text{O}_2 \rightarrow \text{CO}_2$
- $2. \text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- $3. 2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 4\text{CO}_2 + 6\text{H}_2\text{O}$
- $4. \text{C}_2\text{H}_6 + 5\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$
- $5. 2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 4\text{CO}_2 + 6\text{H}_2\text{O}$
- $6. 2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 4\text{CO}_2 + 6\text{H}_2\text{O}$
- $7. 2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 4\text{CO}_2 + 6\text{H}_2\text{O}$
- $8. \text{C}_2\text{H}_6 + 5\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$
- $9. \text{C}_2\text{H}_6 + 5\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$
- $10. 2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 4\text{CO}_2 + 6\text{H}_2\text{O}$

### Part B: Synthesis Reactions

- $1. \text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
- $2. \text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- $3. 2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- $4. 4\text{H} + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- $5. 2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- $6. \text{Fe} + \text{S} \rightarrow \text{FeS}$
- $7. 2\text{Ca} + \text{O}_2 \rightarrow 2\text{CaO}$
- $8. 2\text{H} + \text{S} \rightarrow 2\text{HS}$
- $9. 2\text{H} + \text{S} \rightarrow 2\text{HS}$
- $10. 4\text{H} + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

### Part C: Decomposition Reactions

- $1. 2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
- $2. \text{H}_2\text{CO}_3 \rightarrow \text{H}_2\text{O} + \text{CO}_2$
- $3. 2\text{H}_2\text{CO}_3 \rightarrow 2\text{H}_2\text{O} + 2\text{CO}_2$
- $4. \text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
- $5. 2\text{H}_2\text{CO}_3 \rightarrow 2\text{H}_2\text{O} + 2\text{CO}_2$
- $6. \text{H}_2\text{CO}_3 \rightarrow \text{H}_2\text{O} + \text{CO}_2$
- $7. \text{H}_2\text{CO}_3 \rightarrow \text{H}_2\text{O} + \text{CO}_2$
- $8. \text{H}_2\text{CO}_3 \rightarrow \text{H}_2\text{O} + \text{CO}_2$
- $9. 2\text{H}_2\text{CO}_3 \rightarrow 2\text{H}_2\text{O} + 2\text{CO}_2$
- $10. 2\text{H}_2\text{CO}_3 \rightarrow 2\text{H}_2\text{O} + 2\text{CO}_2$

### Part B: Single Replacement Reactions

- $1. \text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
- $2. \text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$
- $3. \text{Cl}_2 + 2\text{K} \rightarrow 2\text{KCl} + \text{Cl}_2$
- $4. \text{H}_2 + 2\text{HCl} \rightarrow 2\text{H}_2\text{Cl} + \text{H}_2$
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- $6. 2\text{H} + 2\text{HCl} \rightarrow 2\text{H}_2\text{Cl} + 2\text{H}$
- $7. \text{H}_2 + 2\text{HCl} \rightarrow 2\text{H}_2\text{Cl} + \text{H}_2$
- $8. \text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$
- $9. \text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + 2\text{H}_2$
- $10. \text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$

### Part B: Double Replacement Reactions

- $1. \text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{H}_2\text{CO}_3$
- $2. \text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$
- $3. \text{BaCl}_2 + \text{Na}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$
- $4. \text{Pb(NO}_3)_2 + 2\text{KI} \rightarrow \text{PbI}_2 + 2\text{KNO}_3$
- $5. \text{CaCl}_2 + \text{Na}_2\text{CO}_3 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$
- $6. 2\text{HCl} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{SO}_4 + 2\text{HCl}$
- $7. \text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- $8. \text{FeCl}_3 + 3\text{NaOH} \rightarrow \text{Fe(OH)}_3 + 3\text{NaCl}$
- $9. 2\text{HNO}_3 + \text{BaCl}_2 \rightarrow 2\text{HCl} + \text{Ba(NO}_3)_2$
- $10. \text{CuSO}_4 + 2\text{NaOH} \rightarrow \text{Cu(OH)}_2 + \text{Na}_2\text{SO}_4$

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