



Section A

Answers

1. D	6. D	11. A	16. B	21. B	26. B	31. A
2. D	7. C	12. C	17. D	22. D	27. B	32. D
3. A	8. B	13. C	18. A	23. A	28. C	33. C
4. C	9. B	14. B	19. D	24. C	29. D	
5. C	10. C	15. B	20. D	25. B	30. A	

Solutions

1. B

Consider a time period of 1 minute. Volume of water heated by the heater is 900g.

$$P \Delta t = m \cdot c \Delta T$$

$$1800 \times 1000 \times 60 = 900 \times 42000 \cdot (10 - 20)$$

$$\therefore c = 500\text{Jkg}^{-1}\text{K}^{-1}$$

2. D

a. A. The horizontal line represents the change of state. P has higher melting point (melting point) than Q.

b. B. The magnitude of the slope of P in liquid state is smaller than that in solid state. Thus, the specific heat capacity of P in liquid state is larger than that in solid state.

c. C. For a non-constant-time graph, slope $= \frac{\Delta T}{\Delta t} = \frac{T_2 - T_1}{t_2 - t_1}$

As the magnitude of the slope of P in liquid state is smaller than the slope of Q in liquid state, the heat capacity of P in liquid state is greater than that of Q in liquid state.

d. D. The horizontal portion of the lines represents the latent heat of fusion.

The latent heat of fusion of P is smaller than that of Q.

However, since the mass of the two liquids may not be equal,

the specific latent heat of fusion of P may not be smaller than that of Q.

3. B

The specific latent heat of vaporisation is found by $\Delta E = E_f - E_i$

E is the energy supplied to the water and ΔE is the decrease of mass measured by the balance.

A. If water reaches over, ΔE increases, thus ΔE measured becomes smaller.

B. If there is heat loss, ΔE increases (PDR) less water evaporates and ΔE decreases, thus ΔE measured becomes greater.

C. If the temperature does not reach the boiling point, more heat is required to heat up the water to 100°C. Thus, ΔE increases, and ΔE measured becomes greater.

D. If some water vapour condenses and stops heat, ΔE decreases, thus ΔE measured becomes greater.

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Simon Poon

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