

## AP Chemistry Chapter 4 Practice Exam

**D**

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.



How much sulfur (in grams) is produced from 21.5 grams of  $\text{O}_2$ ?

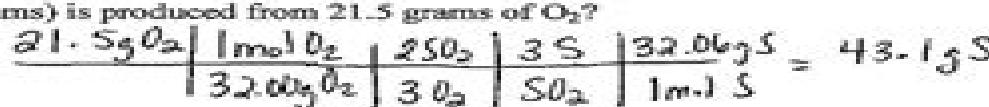
A) 97.0 g

B) 14.4 g

C) 21.5 g

**D) 43.1 g**

E) none of these

**B**

2. Iron is produced from its ore by the reactions:



How many moles of  $\text{O}_2(g)$  are needed to produce 4.6 moles of  $\text{Fe}(s)$ ?

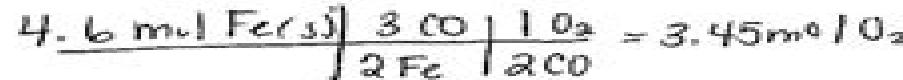
A) 2.3 mole  $\text{O}_2$

**B) 3.5 mole  $\text{O}_2$**

C) 4.6 mole  $\text{O}_2$

D) 6.9 mole  $\text{O}_2$

E) 13.8 mole  $\text{O}_2$

**B**

3. Suppose the reaction  $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$  is carried out starting with 153 g of  $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of  $\text{H}_2\text{SO}_4$ . How much phosphoric acid will be produced?

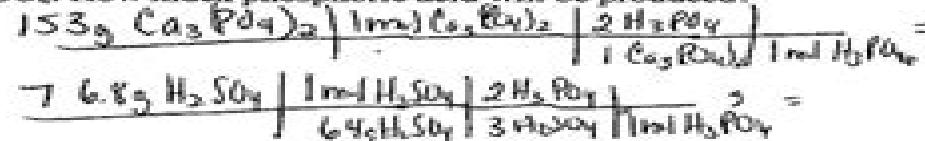
A) 76.7 g

**B) 51.1 g**

C) 229.8 g

D) 115.1 g

E) 96.7 g

**B**

4.  $\text{SO}_2$  reacts with  $\text{H}_2\text{S}$  as follows:



When 7.50 g of  $\text{H}_2\text{S}$  reacts with 12.75 g of  $\text{SO}_2$ , which statement applies?

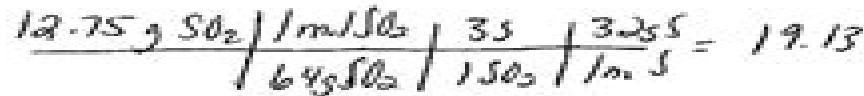
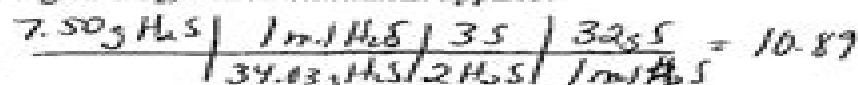
A) 6.38 g of sulfur are formed.

**B) 10.6 g of sulfur are formed.**

C) 0.0216 moles of  $\text{H}_2\text{S}$  remain.

D) 1.13 g of  $\text{H}_2\text{S}$  remain.

E)  $\text{SO}_2$  is the limiting reagent.



# 2013 Ap Chemistry Test Answers With Explanations

**Richard H. Langley, John Moore**

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**The American Catalogue**, 1908

**Journal** New South Wales. Parliament. Legislative Council, 1876

[Catalog](#) Oregon State Agricultural College, 1968

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