


Universal gas constant	R	$= 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$
Avogadro constant	N_A	$= 6.022 \times 10^{23} \text{ mol}^{-1}$
Planck's constant	h	$= 6.626 \times 10^{-34} \text{ J s}$
Velocity of light	c	$= 3 \times 10^8 \text{ m s}^{-1}$

- Which of the following elements has the lowest third ionization energy?
 (1) Mg (2) Ne (3) N (4) P (5) Cl
- Which molecule from the molecules given below consists the highest number of pi (π) bonds?
 (1) H_2SO_4 (2) H_2SO_3 (3) HNO_3 (4) H_3PO_4 (5) HClO_4
- Which statement is true regarding $[\text{Al}(\text{OH})_4]^-$ ion?
 (1) Its central atom hybridization is sp^3 .
 (2) Its total number of lone pairs of electrons is 8.
 (3) It contains d electrons.
 (4) Its number of sigma (σ) bonds is 4.
 (5) Its total number of electrons in the valence shell is 28.
- The IUPAC name of the compound $\text{CH}_3-\text{O}-\underset{\text{O}}{\underset{\parallel}{\text{C}}}-\text{CH}=\underset{\text{CH}_2\text{CH}_3}{\underset{|}{\text{C}}}-\underset{\text{OH}}{\underset{|}{\text{CH}}}-\text{CH}_3$ is
 (1) methyl-3-ethyl-4-hydroxypent-2-enoate
 (2) methyl 3-ethyl-4-hydroxypent-2-enoate
 (3) 3-ethyl-1-methoxy-1-oxopent-3-en-4-ol
 (4) 3-ethyl-5-methoxy-5-oxopent-3-en-2-ol
 (5) methyl 3-ethyl-2-hydroxypent-3-enoate
- The production process which produces a gas as a by-product giving the highest contribution to global warming is.
 (1) Soap production (2) Nitric acid production (3) Iron production
 (4) Sulphuric acid production (5) Bio - diesel production
- Which one of the following compounds undergoes self condensation in basic condition?
 (1) $\text{CH}_3-\underset{\text{CH}_3}{\underset{|}{\overset{\text{OH}}{\text{C}}}}-\underset{\text{H}}{\underset{|}{\text{C}}}=\text{O}$ (2) $\text{CH}_3-\underset{\text{CH}_3}{\underset{|}{\overset{\text{CH}_3}{\text{C}}}}-\underset{\text{H}}{\underset{|}{\text{C}}}=\text{O}$ (3) $\text{CH}_3-\underset{\text{H}}{\underset{|}{\text{C}}}=\text{O}$
 (4)  (5) $\text{H}-\underset{\text{H}}{\underset{|}{\text{C}}}=\text{O}$
- Consider the following reaction at 25°C .
 $\text{CO}_2(\text{g}) + 2\text{NH}_3(\text{g}) \longrightarrow \text{CO}(\text{NH}_2)_2(\text{s}) + \text{H}_2\text{O}(\text{l}) ; \Delta H^\circ = -134 \text{ kJ mol}^{-1}$
 Which of the following is correct regarding this reaction?
 (1) ΔS° is always a negative value for the reaction.
 (2) ΔH° increases with the temperature.
 (3) The decrease in the entropy change can determine the spontaneity of the reaction.
 (4) The reaction is spontaneous at all temperatures.
 (5) At high temperatures the value of ΔG° becomes more negative.

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Yin Toon Tan



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