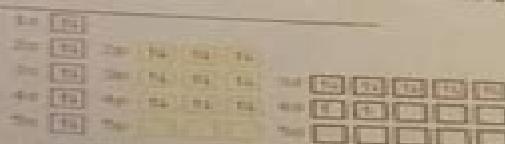


Unit 2.10 Electron Configuration Practice Worksheet-30pts (2 pts each)
 You may use a periodic table and any notes that you have. If you need help, refer to the recorded ClassConnect session or this Khan academy video:

<https://www.khanacademy.org/science/8th-chemistry-basics/2ccf869c74cd302/atomic-structure-and-properties/x/2ccf869c74cd302/atomic-structure-and-electron-configuration/introduction-to-electron-configurations>

1. List the suborbitals (letters) found at the 3⁻ energy level: _____
2. How many total electrons can a "p" suborbital hold? _____
3. How many total electrons can the 2⁻ energy level hold? (don't forget to take into account each suborbital found there). _____
4. Identify the following element based on its orbital filling diagram:



5. If Mg loses 2 electrons and becomes Mg²⁺, how many total electrons does the ion now have? _____

Use the FULL electron configuration for the following elements. Note that some can have more than one correct answer.

6. Sodium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^5$
7. Iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
8. Bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^4 4p^5$
9. Barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^10 4p^6 5s^2 4f^10 5p^6 6s^2$
10. Cu²⁺ $1s^2 2s^2 2p^6 3s^2 3p^6 3d^9$

Determine the elements for the following electron configurations.

- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ Sulfur
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^5 4p^5$ Publith
- 13) $[Kr] 5s^2 4d^1 5p^3$ Antimony
- 14) $[Xe] 6s^2 4f^1 5d^1$ Chlorine
- 15) $[Rn] 7s^2 5f^1$ Platinum

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