

You may use a periodic table and any notes that you have. If you need help, refer to the recorded ClassConnect session or this Khan academy video:
<https://www.khanacademy.org/science/chemistry/electron-configuration/a/electron-configuration-v1>

1. List the suborbitals (d-orbitals)

- How many total electrons can a "p" suborbital hold?
- How many total electrons can the 2^- energy level hold? (don't forget to take into account each suborbital found there).
- Identify the following element based on its orbital filling diagram:

[illegible]

5. If Mg loses 2 electrons and becomes Mg^{2+} , how many total electrons does the ion now have?

Write the FILL electron configuration for the following elements. Fill each box with the appropriate number of electrons. (diagram.)

6. Sodium $1s^2 2s^2 2p^6 3s^1$
7. Iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
8. Bromine $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^5 4p^5$
9. Barium $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^0 4p^6 5s^2 4d^0 5p^6 6s^2$
10. Cu^{2+}

Determine the elements for the following electron configurations.

- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ Sulfur
12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ Rubidium
13) $[Kr] 5s^2 4d^5 5p^3$ Antimony
14) $[Xe] 6s^2 4f^1 5d^5$ Osmium
15) $[Rn] 7s^2 5f^{11}$ Einsteinium

Apex Chemistry Answers Unit 3 Lesson 5

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