

ACE Practice Test

Chemistry, Sixth Edition

Chapter 13-1 : Chapter 13: Chemical Equilibrium

1. When chemical equilibrium is achieved.
 - A. reactants are no longer being converted to products and products are no longer being converted to reactants
 - B. only reactants continue to be converted to products
 - C. only products continue to be converted to reactants
 - D. the amount of each reactant and product remains constant with time

2. Hydrogen fluoride, HF, decomposes to produce hydrogen gas and fluorine gas. If 80.0 g of HF were initially placed in a reaction vessel, calculate the amount of HF that was left if equilibrium was reached when 2.0 g of hydrogen had been produced.
 - A. 20 g
 - B. 25 g
 - C. 34 g
 - D. 40 g
 - E. 60 g

3. SO reacts with O₂ to produce SO₂. If 168.0 g of SO were initially placed in a reaction vessel containing excess oxygen gas, calculate the number of moles of SO that was left if equilibrium was reached when 96.0 g of SO₂ had been produced.
 - A. 0.75
 - B. 1.5
 - C. 2.0
 - D. 3.0
 - E. 4.0

Ace Practice Test Chemistry

**Karen Lurie, Magda Pecsene, Adam
Robinson**

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