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## ACE Practice Test

Chemistry, Sixth Edition

### Chapter 13-1 : Chapter 13: Chemical Equilibrium

1. When chemical equilibrium is achieved,
  - A. reactants are no longer being converted to products and products are no longer being converted to reactants
  - B. only reactants continue to be converted to products
  - C. only products continue to be converted to reactants
  - D. the amount of each reactant and product remains constant with time
  
2. Hydrogen fluoride, HF, decomposes to produce hydrogen gas and fluorine gas. If 80.0 g of HF were initially placed in a reaction vessel, calculate the amount of HF that was left if equilibrium was reached when 2.0 g of hydrogen had been produced.
  - A. 20 g
  - B. 25 g
  - C. 34 g
  - D. 40 g
  - E. 60 g
  
3. SO reacts with O<sub>2</sub> to produce SO<sub>2</sub>. If 168.0 g of SO were initially placed in a reaction vessel containing excess oxygen gas, calculate the number of moles of SO that was left if equilibrium was reached when 96.0 g of SO<sub>2</sub> had been produced.
  - A. 0.75
  - B. 1.5
  - C. 2.0
  - D. 3.0
  - E. 4.0

# Ace Practice Test Chemistry

**Karen Lurie, Magda Pecsénye, Adam  
Robinson**



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