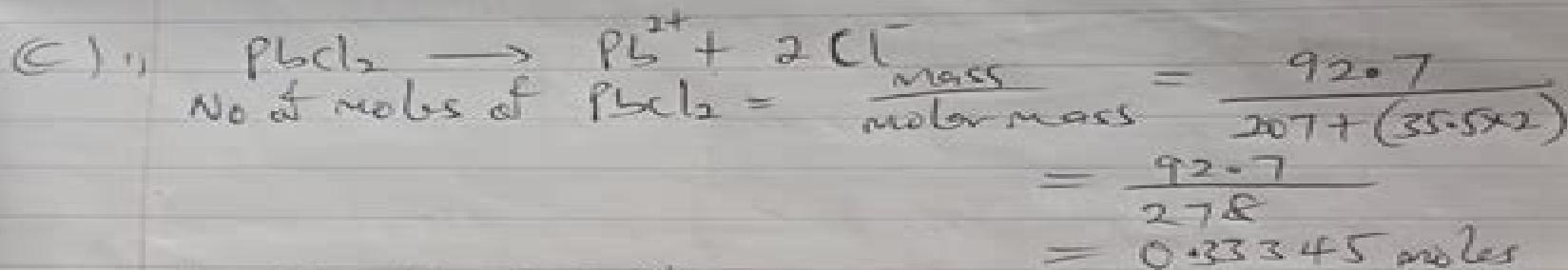


Question 2 (contd)

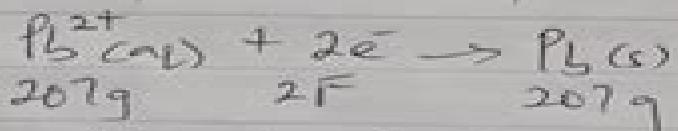
⑤ (a) The system is an exothermic one, when it is cooled, the equilibrium will adjust to annul the cooling effect by releasing more heat. And this favours the forward reaction.



No. of moles of Pb^{2+} ionized = 0.33345 moles

mass of lead present = no. of moles \times molar mass
 $= 0.33345 \times 207$
 $= 69.0 \text{ grams}$

(v) Quantity of electricity used = $It = I \times 4.8 \times 3600$
 $= (17280 I) \text{ C}$



2×96500 will liberate 207 grams of Pb

17280 I will liberate $17280 I \times 207 = 207$

$$I = \frac{2 \times 96500 \times 207}{17280 \times 207} = \frac{2 \times 96500}{17280} \text{ Amperes}$$

$$= 11.17 \text{ Amperes}$$

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