

# CHEMISTRY MULTIPLE CHOICE QUESTIONS

## A. Atoms, molecules and stoichiometry

2002 -2019



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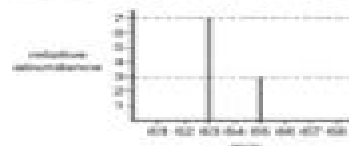
As a simplification, an adult human can be considered to have a daily diet of 1.80 kg of carbohydrate (empirical formula  $C_6H_{12}O_6$ ).

Which mass of carbon dioxide does a person produce each day if all the carbohydrate eaten is digested and oxidised?

- A. 0.907 kg    B. 0.900 kg    C. 1.30 kg    D. 2.54 kg

(2002 NA/12)

The diagram shows the mass spectrum of a sample of naturally-occurring copper.



What is the relative atomic mass of this copper?

- A. 63.2    B. 63.5    C. 63.6    D. 64.0

(2002 NA/10)

Which pairs of compounds have the same empirical formula?

- ethane and ethene
- ethane and cyclohexane
- cyclohexane and oct-1-ene

(2002 MC/31)

A mixture of 1.00 g of methane and 1.00 g of ethane was sparked with an excess of oxygen. After cooling to room temperature, the residual gas was passed through aqueous potassium hydroxide.

Which compounds have the empirical formula  $CH_4O$ ?

- methanol
- ethanoic acid
- methyl methanolate

(2002 ON/31)

What is the number of molecules in 500 cm<sup>3</sup> of oxygen under room conditions?

- A.  $1.25 \times 10^{23}$     B.  $1.25 \times 10^{24}$     C.  $3.0 \times 10^{23}$     D.  $3.0 \times 10^{24}$

(2008 MC/12)

In the preparation of soft margarine, glycerol trioleate reacts



is fully hydrogenated so that, on average, one of its side-chains is converted into the  $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2(\text{CH}_2)_7\text{CO}_2$  moiety and two side-chains are converted into the  $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2(\text{CH}_2)_7\text{CO}_2$  moiety.

How many moles of hydrogen are required to convert one mole of glycerol trioleate into the soft margarine?

- A. 4    B. 5    C. 6    D. 8

(2008 MC/22)

Use of the Data Booklet is relevant to this question.

Additional classes are about very small amounts of water which dissolve in 1000 g of

# 2012 Hsc Chemistry Answers Multiple Choice

**Yin Toon Tan**



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