

Multiple Choices: Entropy

1. In which of the following does the entropy decrease?

- A) $\text{NaCl}_{(s)} \rightarrow \text{Na}^{+}_{(aq)} + \text{Cl}^{-}_{(aq)}$
- B) $4\text{Na}_{(s)} + \text{S}^{2-}_{(aq)} \rightarrow 4\text{Na}^{+}_{(aq)} + \text{S}^{2-}_{(aq)}$
- C) $2\text{NaNO}_3_{(s)} \rightarrow \text{Na}_2\text{O}_{(s)} + \text{O}_2(g) + \text{NaNO}_2$
- D) $\text{CaCO}_3 + 2\text{HCl}_{(aq)} \rightarrow \text{CaCl}_{(aq)} + \text{CO}_2(g) + \text{H}_2\text{O}_{(l)}$

ANS: B

1. The normal boiling point of methanol is -91.7°C and the molar enthalpy of vaporization is 73.0 kJ/mol . The value of ΔS when 2.15 mol of CH_3OH vaporizes at -91.7°C is _____.

- A) 0.457
- B) 5.21×10^3
- C) 457
- D) 1.39×10^3
- E) 2.39

ANS: C

1. Which one of the following reactions has a positive entropy change?

- A) $\text{NaCl}_{(s)} \rightarrow \text{Na}^{+}_{(aq)} + \text{Cl}^{-}_{(aq)}$
- B) $\text{S}_{(s)} + 2\text{NH}_3_{(g)} \rightarrow \text{NH}_4\text{NH}_3_{(s)}$
- C) $2\text{H}_2\text{O}_{(g)} + \text{O}_{(g)} \rightarrow 2\text{H}_2\text{O}_{(l)}$
- D) $\text{Na}_{(s)} + 3\text{H}_{(g)} \rightarrow 2\text{NaH}_{(s)}$
- E) $2\text{NaH}_2\text{O}_{(g)} \rightarrow 2\text{Na}_{(s)} + 4\text{H}_2\text{O}_{(g)} + \text{O}_{(g)}$

ANS: B

1. For a certain process at 127°C , $\Delta G = -16.00 \text{ kJ}$ and $\Delta H = -1700 \text{ kJ}$. What is the entropy change for this process at this temperature? Express your answer in the form, $\Delta S = \text{_____ J/K}$.

- A) -4.3 J/K
- B) -6.3 J/K
- C) -2.0 J/K
- D) -2.0 J/K
- E) -0.1 J/K

ANS: C

1. Evaluate ΔS° for the reaction below at 295°C .



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Richard H. Langley, John Moore

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