

## Multiple Choices: Entropy

1. In which of the following does the entropy decrease?

- A)  $\text{NaCl(s)} \rightarrow \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$   
 B)  $4\text{NO(g)} + 6\text{H}_2\text{O(g)} \rightarrow 4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g})$   
 C)  $2\text{NaHCO}_3(\text{s}) \rightarrow \text{Na}_2\text{CO}_3(\text{s}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O(g)}$   
 D)  $\text{CaCO}_3(\text{s}) + 2\text{HCl(aq)} \rightarrow \text{CaCl}_2(\text{aq}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O(l)}$

ANS: B

1. The normal boiling point of methanol is  $64.7^\circ\text{C}$  and the molar enthalpy of vaporization is  $71.6\text{ kJ/mol}$ . The value of  $\Delta S$  when  $2.15\text{ mol}$  of  $\text{CH}_3\text{OH(l)}$  vaporizes at  $64.7^\circ\text{C}$  is \_\_\_\_\_  $\text{J/K}$ .

- A)  $0.457$                       B)  $521 \times 10^3$   
 C)  $457$                       D)  $2.39 \times 10^3$   
 E)  $2.39$

ANS: C

1. Which one of the following reactions has a positive entropy change?

- A)  $\text{H}_2\text{O(g)} \rightarrow \text{H}_2\text{O(l)}$   
 B)  $\text{BF}_3(\text{g}) + \text{NH}_3(\text{g}) \rightarrow \text{F}_3\text{BNH}_3(\text{g})$   
 C)  $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_3(\text{g})$   
 D)  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$   
 E)  $2\text{NH}_4\text{NO}_3(\text{s}) \rightarrow 2\text{N}_2(\text{g}) + 4\text{H}_2\text{O(g)} + \text{O}_2(\text{g})$

ANS: E

1. For a certain process at  $127^\circ\text{C}$ ,  $\Delta G = -16.20\text{ kJ}$  and  $\Delta H = -17.0\text{ kJ}$ . What is the entropy change for this process at this temperature? Express your answer in the form,  $\Delta S = \frac{\text{J/K}}$ .

- A)  $-6.3\text{ J/K}$                       B)  $+6.3\text{ J/K}$                       C)  $-2.0\text{ J/K}$   
 D)  $+2.0\text{ J/K}$                       E)  $-0.1\text{ J/K}$

ANS: C

1. Evaluate  $\Delta S^\circ$  for the reaction below at  $25^\circ\text{C}$ .



	$\text{CH}_4(\text{g})$	$+ 2\text{Cl}_2(\text{g})$	$\rightarrow$	$\text{CCl}_4(\text{l})$	$+ 2\text{H}_2(\text{g})$
$\Delta G^\circ_f \text{ (kJ/mol)}$	$-50.5$	$0$		$-63.4$	$0$
$\Delta G^\circ_f \text{ (kJ/mol)}$	$-16.7$	$0$		$-67.7$	$0$

- A)  $-360\text{ J/K}$                       B)  $-66.9\text{ J/K}$                       C)  $-155\text{ J/K}$   
 D)  $-467\text{ J/K}$                       E)  $-387\text{ J/K}$

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**Richard H. Langley, John Moore**



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