

TEST YOUR SELF

TEST YOURSELF 2.1 (THEORIES AND EXPERIMENTS RELATED TO STRUCTURE OF ATOMS)

i) Do you know any element having no neutrons in its atoms?

Ans: Yes the isotope of hydrogen (Protium, ${}^1_1\text{H}$) has no neutron. It has one proton in its nucleus and one electron revolving around it.

ii) Who discovered an electron, a proton and a neutron?

Ans: a. In 1886, Goldstein discovered positively charged particles called protons.

b. In 1897, J.J. Thomson discovered the negatively charged particles called electrons.

c. In 1932, James Chadwick discovered the neutral particles called neutrons.

iii) How does electron differ from a neutron?

Ans: Difference between electron and neutron.

Electron	Neutron
i. Electron is the negatively charged particle.	i. Neutron is the neutral particle.
ii. Electron revolves around the nucleus.	ii. Neutron is present in the nucleus.
iii. Mass of electron is 5.486×10^{-4} g or 9.106×10^{-28} g.	iii. Mass of neutron is 1.0087 amu or 1.674×10^{-24} g.

iv) Explain how anode rays are formed from the gas taken in the discharge tube?

Ans: Anode rays do not originate from the anode. In fact these rays are produced when the cathode rays or electrons collide with the residual gas molecules present in the discharge tube and ionize them as follows:



TEST YOURSELF 2.2 RUTHERFORD'S ATOMIC MODEL AND BOHR'S ATOMIC MODEL

i) How was it proved that the whole mass of an atom is located at its centre?

Ans: Rutherford bombarded alpha particles on a 0.00004 cm thick gold foil. Almost all the particles passed through the foil undeflected, only few bounced back. The complete rebound of few alpha particles showed that the nucleus is an extremely small positively charged part. It is situated at the center of an atom and it carries nearly the whole mass of an atom.

ii) How was it shown that atomic nuclei are positively charged?

Ans: Alpha particles are helium nuclei (He^{2+}) i.e. doubly positively charged particles. The deflection and rebounding of a few particles showed that atomic nuclei (centre of an atom) are positively charged.

iii) Name the particles which determine the mass of an atom.

Ans: Protons and neutrons present in the nucleus determine the mass of an atom. The sum of number of protons and neutrons is equal to mass of an atom.

iv) What is the classical theory of radiation? How does it differ from quantum theory?

Ans:

Classical theory of radiation	Quantum theory of radiation
According to classical theory of radiation electrons being the charged particles should release or emit energy continuously and they should ultimately fall into the nucleus.	According to quantum theory of radiation, energy is not emitted or absorbed continuously but it is emitted or absorbed in the form of small energy packets or bundles known as quantum or

Quiz Yourself Chapter 8 Introductcion To Chemistry

Steven L. Hanft



Quiz Yourself Chapter 8 Introductcion To Chemistry:

Student's Guide to Introduction to Chemical Principles by Edward I. Peters, 2d Ed Peter P. Berlow, 1978

Aggregate Base and Surfacing Inspection , 1986

Forensic Science Jay A. Siegel, Kathy Mirakovits, 2006-09-07

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Exam Educart, Sir Tarun Rupani, 2024-06-28 Study Guide for Chemical Principles, Fourth Edition, by Dickerson, Gray, Darensbourg, and Darensbourg Patricia L. Samuel, 1984 **Basic Chemistry** G. William Daub, William S. Seese, 1996

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and includes reviews of all test areas from writing skills to science *How to Prepare for the GED with CD-ROM* Murray Rockowitz, Samuel C. Brownstein, Ira K. Wolf, Max Peters, 2002-01-30 Updated to reflect the recent major changes in the high school equivalency exam this manual presents a full length diagnostic test with answer keys answer analyses and self appraisal charts The diagnostic is followed by reviews of all GED test subjects which emphasize proficiency in high school level math reading comprehension and writing skill as well as knowledge of social studies arts and literature Two full length practice exams modeled after the new GED exam complete this manual They come with all questions answered and explained The CD ROM packaged with this book offers a computerized practice GED with automatic scoring plus additional subject review material **Student-active Science** Ann P. McNeal, 1997 This volume emerged from an NSF sponsored conference on Inquiry Approaches to Science Teaching held at Hampshire College in June 1996 STUDENT ACTIVE SCIENCE emphasizes that experiencing the process of science is central to the learning of science This book is a collection of articles ideas and models for science education reform and is the result of collaboration between instructors frustrated with the traditional approach to teaching You ll find models and ideas that promote critical thinking and hands on science in the classroom as well as commentary from school wide department wide and individual reform efforts Edexcel Modular , 2010-02-22 Collins New GCSE Maths Edexcel Modular Teacher s Pack Foundation 1 contains everything you need to deliver effective lessons in mathematics with confidence for students working at Grades G to C Fully matched to Edexcel s new GCSE Maths Modular specification these teacher resources offer well differentiated lesson plans and additional support Nature , 1895 The School World , 1899

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