

Chemistry 2nd Semester Final Exam Review

Chemical Bonds

1. Give a physical description of how the atoms and molecules are arranged in solids, liquids, and gases.
A: In a liquid, the forces between the molecules are weak, so the molecules are free to move around. In a solid the forces between the molecules (intermolecular forces) are stronger so they are held in place.
2. Use the information from the previous question to describe the shape and volume for each phase (state) of matter.
*A: Solids have a repeating geometric pattern (lattice structure) in which the particles vibrate in place. They have a definite shape and a definite volume.
Particles within a liquid slide past each and are close, but are not held in place. They have a definite volume, but no definite shape.
Gases particles move chaotically bouncing off one another. They have neither a definite shape nor a definite volume.*
3. List the three phases of matter in order of increasing intermolecular attractions.
A: Gas, liquid, solid.
4. Why do the atoms and molecules in liquids move in a random pattern relative to one another instead of being in a solid form?
A: Their kinetic energy (temperature) is high enough to overcome the attractive forces that might otherwise hold them in a repeating pattern. The attractive forces aren't strong enough to hold them together at that temperature.

Conservation of Matter and Stoichiometry

5. The atomic mass unit and the mole are defined using what isotope? *A: Carbon-12*
6. Balance the following equations:
 - a. $\text{Al}_2\text{O}_3 + \text{Cl}_2 + \text{C} \rightarrow \text{AlCl}_3 + \text{CO}$
A: $\text{Al}_2\text{O}_3 + 3\text{Cl}_2 + 3\text{C} \rightarrow 2\text{AlCl}_3 + 3\text{CO}$
 - b. $\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
A: $2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$
 - c. $\text{P} + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$
A: $4\text{P} + 5\text{O}_2 \rightarrow 2\text{P}_2\text{O}_5$
7. What isotope is used to define one mole? *A: Carbon-12*
8. How many particles (atoms or molecules) are in one mole? *A: 6.02×10^{23}*
9. Find the molar mass for the following chemical formulas:
 - a. ZnF_2 *A: 103.377 g/mol*
 - b. $\text{Al}_2(\text{SO}_4)_3$ *A: 342.151 g/mol*
 - c. NH_4IO_3 *A: 192.941 g/mol*
10. How many moles are in 4.37 kg of NaOH? *A: 109.3 mol*
11. How many atoms are in 2075 g of He? *A: 3.12×10^{26} atoms*
12. What volume will 92.31 g of CO_2 occupy at STP? *A: 46.985 L*
13. How many grams of sodium hydroxide are needed to react with 6.23 g of barium bromide?
 $\text{NaOH} + \text{BaBr}_2 \rightarrow \text{NaBr} + \text{Ba(OH)}_2$
A: 1.68 g NaOH
14. In the reaction, $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$, if 100.0 g of magnesium reacts with 50.0 g of oxygen, what mass of product is produced?
A: 126 g MgO

Gases and Their Properties

15. What is the kinetic molecular theory (KMT)?
A: The Kinetic Molecular Theory is used to explain the behavior of gases and is based upon the following postulates:
 - Gases are composed of a many particles that behave like hard spherical objects in a state of constant, random motion.
 - These particles move in a straight line until they collide with another particle or the walls of the container.
 - These particles are much smaller than the distance between particles, therefore the volume of a gas is mostly empty space and the volume of the gas molecule themselves is negligible.
 - There is no force of attraction between gas particles or between the particles and the walls of the container.
 - Collisions between gas particles or collisions with the walls of the container are elastic. That is, none of the energy of the gas particle is lost in a collision.

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Cornell University Courses of Study Cornell University,2003 **The School Review** ,1914 Source Book for Chemistry Teachers William Thomas Lippincott,1981 Summer School Number Southwestern at Memphis,1933 **Lucky Break** Howard Junker,1999 In Lucky Break Howard Junker introduces you to nineteen contemporary authors who describe their struggles and good fortune in transforming themselves into writers For some the idea of being a writer seemed to run in the family For others writing was a way of breaking with their families for still others writing was the only way to find their own voice to preserve what is uniquely theirs Collected from the pages of ZYZZYVA the West Coast literary journal these are very personal essays in which nevertheless other writers will discover something of their own stories and journeys Among the authors featured in Lucky Break are Sheila Ballantyne Opal Palmer Adisa Philip Levine Jewelle Gomez Justin Chin David Rains Wallace and Sallie Tisdale Becoming a writer is a mysterious momentous process The writers in this collection suggest that beyond talent and hard work and a lucky break there is more hard work the task of being a writer They offer accounts not of the craft of writing but of the many strands that must be woven together to create an identity as a writer *Current Index to Journals in Education* ,1974 Catalog Issue University of Colorado,1974 *Zyzzzyva* ,1999 **Resources in Education** ,1978 Paperbound Books in Print ,1966 *American Journal of Physics* ,1994 **Science Books & Films** ,1980 Computer Aided Training in Science and Technology E. Oñate,1990 **Current Index to Journals in Education, Semi-Annual Cumulation, July-December, 1976** ,1976-10

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