

Guided Practice - Acids & Bases**Properties of Acids and Bases****1. Match the following terms with their descriptions:**

- a. Autoionized Acid _____ the substance formed when a proton is added to _____ a base
- b. Autoionized Base _____ the remaining substance when a proton is taken from an acid
- c. Conjugate Acid _____ a substance that produces OH^- in solution
- d. Conjugate Base _____ a substance producing H^+ in solution
- e. Bronsted-Lowry Acid _____ a proton acceptor
- f. Bronsted-Lowry Base _____ a proton donor

2. Which of the following represent conjugate acid-base pairs? Label each species as a conjugate acid or a conjugate base. For those pairs that are NOT conjugates, write the correct conjugate acid or base for each species in the pair, labeling each species correctly.

- a. HSO_4^- , SO_4^{2-}
- b. HBr , BrO^-
- c. H_2PO_4^- , PO_4^{3-}
- d. H_2O^{2-} , HO^{2-}

3. Complete the following table regarding acid strength.

Property/Condition	Strong Acid	Weak Acid
Conjugate base		
Species present in solution		
Relative strengths of conjugate base compared to water		
Equations following acids in order from weakest		

Answer Key To Describing Acids And Bases

**BioPharmaceutical Technology Center
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Answer Key To Describing Acids And Bases:

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Magill's Medical Guide Anne Chang,2005

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