

17 • Acid-Base Equilibria

1. What is the $[H^+]$ when $[OH^-] = 8.1 \times 10^{-3}$?

- $8.1 \times 10^{-3} M$
- $1.0 \times 10^{-7} M$
- $1.2 \times 10^{-20} M$ $(H^+) = \frac{K_w}{[OH^-]} = \frac{1.0 \times 10^{-14}}{8.1 \times 10^{-3}} =$
- $3.6 \times 10^{-6} M$
- $8.1 \times 10^{-3} M$

2. What is the $[H^+]$ when $[OH^-] = 3.3 \times 10^{-7}$?

- $3.0 \times 10^{-6} M$
- $1.0 \times 10^{-7} M$
- $3.3 \times 10^{-7} M$ $(H^+) = \frac{K_w}{[OH^-]} = \frac{1.0 \times 10^{-14}}{3.3 \times 10^{-7}} =$
- $6.6 \times 10^{-8} M$
- $3.3 \times 10^{-7} M$

3. What is the $[H^+]$ in a $0.0025 M$ HCl solution?

- $1.0 \times 10^{-3} M$
- $4.0 \times 10^{-12} M$
- $2.5 \times 10^{-3} M$ $(H^+) = \frac{K_w}{[OH^-]} = \frac{1.0 \times 10^{-14}}{0.0025} = 2.5 \times 10^{-12} M$
- $3.6 \times 10^{-5} M$

4. What is the $[OH^-]$ in a $0.0050 M$ HCl solution?

- $5.0 \times 10^{-3} M$
- $1.0 M$
- $1.0 \times 10^{-7} M$ $(H^+) = \frac{K_w}{[OH^-]} = \frac{1.0 \times 10^{-14}}{0.0050} =$
- $6.6 \times 10^{-9} M$

5. A solution in which $[H^+] = 10^{-8}$ has a pH of B and is basic.

- 8, acidic
- 6, basic
- 6, basic
- 8, neutral
- 8, basic

6. What is the pH of a $0.00030 M$ HNO_3 solution?

- 8.11
- 3.00
- 3.52 $(H^+) = 3.4 \times 10^{-4}$
- 4.48
- none of these

7. What is the pH of a $0.0060 M$ KOH solution?

- 5.12
- 2.22
- 11.78 $(OH^-) = 6.0 \times 10^{-3}$
- 8.88
- 7.00

post = 2.22
pH = 14 - pOH =

8. A sample of lemon juice is found to have a pH of 2.55. What is the H^+ concentration of the juice? $(H^+) = 10^{-14} = 10^{-2.55}$

- $0.0035 M$
- $0.0028 M$
- $355 M$
- $0.0080 M$

9. A sample of milk is found to have a pH of 6.60. What is the OH^- concentration of the milk?

- $2.5 \times 10^{-21} M$
- $1.0 \times 10^{-7} M$
- $5.0 \times 10^{-7} M$ $(H^+) = 10^{-6.60} = 2.5 \times 10^{-7}$
- $4.0 \times 10^{-8} M$
- $2.5 \times 10^{-7} M$

10. What is the concentration of Cl^- in a $0.60 M$ solution of $HOCl$? $K_a = 3.1 \times 10^{-8}$

- $1.8 \times 10^{-4} M$
- $7.1 \times 10^{-11} M$
- $0.40 M$ $(OH^-) = (H^+) = Ka \cdot x = x^2$
- $1.4 \times 10^{-4} M$
- $1.1 \times 10^{-4} M$

see scratch paper

11. What is the pH of a $0.020 M$ solution of hydrochloric acid, a diprotic acid? $K_{a1} = 1.1 \times 10^{-7}$ $K_{a2} = 1.0 \times 10^{-14}$ $\text{dissociation is significant}$

- 7.00
- 9.67
- 3.05
- 4.33
- 7.84

see scratch paper

12. What is the concentration of CO_3^{2-} in a $0.010 M$ solution of carbonic acid? The relevant equilibria are:

$$H_2CO_3 \rightleftharpoons H^+ + HCO_3^- \quad K_{a1} = 4.3 \times 10^{-7}$$

$$HCO_3^- \rightleftharpoons H^+ + CO_3^{2-} \quad K_{a2} = 5.6 \times 10^{-11}$$

- $6.6 \times 10^{-5} M$
- $5.6 \times 10^{-11} M$
- $6.7 \times 10^{-11} M$
- $7.5 \times 10^{-7} M$
- $7.9 \times 10^{-7} M$

from the 1st dissociation, $[H^+] = [HCO_3^-]$
the 2nd dissociation doesn't change this much.
 $K_{a2} = \frac{[H^+][CO_3^{2-}]}{[HCO_3^-]}$ but $[H^+] = [HCO_3^-]$ and $[HCO_3^-] \text{ small}$

so $K_{a2} = [CO_3^{2-}] = 5.6 \times 10^{-11}$

Answer Key To Acids Bases And Solutions

**Donald Voet, Judith G. Voet, Charlotte
W. Pratt**

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Analytical Chemistry Bryan M. Ham, Aihui MaHam, 2015-10-26. A comprehensive study of analytical chemistry providing the basics of analytical chemistry and introductions to the laboratory. Covers the basics of a chemistry lab including lab safety, glassware and common instrumentation. Covers fundamentals of analytical techniques such as wet chemistry, instrumental analyses, spectroscopy, chromatography, FTIR, NMR, XRF, XRD, HPLC, GC, MS, Capillary Electrophoresis and proteomics. Includes ChemTech, an interactive program that contains lesson exercises, useful calculators and an interactive periodic table. Details Laboratory Information Management System, a program used to log in samples, input data, search samples, approve samples and print reports and certificates of analysis.

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A. G. Whittaker, A. Mount, M. Heal, 2000-09-21. This book provides concise yet comprehensive coverage of physical chemistry.

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Raymond Chang, 1988. **Journal of Solution Chemistry**, 1993. **Acids, Bases, and Solutions** Elizabeth R. Kellerman, 1999-09-01.

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