

# 17 • Acid-Base Equilibria

- What is the  $[H^+]$  when  $[OH^-] = 8.1 \times 10^{-9}$ ?  
 a)  $8.1 \times 10^{-5} M$  d)  $3.6 \times 10^{-6} M$   
 b)  $1.0 \times 10^{-7} M$  e)  $8.1 \times 10^{-2} M$   
 c)  $1.2 \times 10^{-10} M$   $[H^+] = \frac{K_w}{[OH^-]} = \frac{1 \times 10^{-14}}{8.1 \times 10^{-9}} = 1.2 \times 10^{-6}$
- What is the  $[H^+]$  when  $[OH^-] = 3.3 \times 10^{-9}$ ?  
 a)  $3.0 \times 10^{-6} M$  d)  $6.6 \times 10^{-5} M$   
 b)  $1.0 \times 10^{-7} M$  e)  $3.3 \times 10^{-9} M$   
 c)  $3.3 \times 10^{-5} M$   $[H^+] = \frac{K_w}{[OH^-]} = \frac{1 \times 10^{-14}}{3.3 \times 10^{-9}} = 3.0 \times 10^{-6}$
- What is the  $[H^+]$  in a 0.0025 M HCl solution?  
 a)  $1.0 \times 10^{-3} M$  d)  $3.6 \times 10^{-5} M$   
 b)  $4.0 \times 10^{-12} M$  e) need more info  
 c)  $2.5 \times 10^{-3} M$  HCl strong acid  
 $[H^+] = .0025 = 2.5 \times 10^{-3}$
- What is the  $[OH^-]$  in a 0.0050 M HCl solution?  
 a)  $5.0 \times 10^{-3} M$  d)  $6.6 \times 10^{-5} M$   
 b) 1.0 M e)  $2.0 \times 10^{-12} M$   
 c)  $1.0 \times 10^{-7} M$   $[H^+] = .0050$   
 $[OH^-] = \frac{K_w}{[H^+]} = \frac{1 \times 10^{-14}}{.0050} = 2.0 \times 10^{-12}$
- A solution in which  $[H^+] = 10^{-8}$  has a pH of 8 and is basic.  
 a) 8, acidic d) -8, neutral  
 b) 6, basic e) 8, basic  
 c) -6, basic
- What is the pH of a 0.00030 M  $HNO_3$  solution?  
 a) 8.11 d) 4.48  
 b) 3.00 e) none of these  
 c) 3.52  $[H^+] = 3 \times 10^{-4}$   
 $pH = -\log [H^+] = 3.52$
- What is the pH of a 0.0060 M KOH solution?  
 a) 5.12 d) 8.88  
 b) 2.22 e) 7.00  
 c) 11.78  $[OH^-] = 6.0 \times 10^{-3}$   
 $pOH = 2.22$   
 $pH = 14 - pOH = 11.78$

## PRACTICE TEST

- A sample of lemon juice is found to have a pH of 2.55. What is the  $H^+$  concentration of the juice?  
 $[H^+] = 10^{-pH} = 10^{-2.55}$   
 a) 0.0035 M d) 0.0080 M  
 b) 0.0028 M e) 355 M  
 c) 11.6 M
- A sample of milk is found to have a pH of 6.60. What is the  $OH^-$  concentration of the milk?  
 a)  $2.5 \times 10^{-21} M$  d)  $4.0 \times 10^{-8} M$   
 b)  $1.0 \times 10^{-7} M$  e)  $2.5 \times 10^{-7} M$   
 c)  $5.0 \times 10^{-7} M$   $[H^+] = 10^{-6.60} = 2.5 \times 10^{-7}$   
 $[OH^-] = \frac{K_w}{[H^+]}$
- What is the concentration of OCl<sup>-</sup> in a 0.60 M solution of HOCl?  $K_a = 3.1 \times 10^{-8}$ .  
 a)  $1.8 \times 10^{-4} M$  d)  $1.4 \times 10^{-4} M$   
 b)  $7.1 \times 10^{-11} M$  e)  $1.1 \times 10^{-4} M$   
 c) 0.40 M  $[OCl^-] = [H^+] \bullet \frac{K_a \bullet .6}{x^2}$   
~~see scratch paper~~
- What is the pH of a 0.020 M solution of hydrosulfuric acid, a diprotic acid?  
 $K_{a1} = 1.1 \times 10^{-7}$   $K_{a2} = 1.0 \times 10^{-14}$  only 1st dissociation is signif.  
 a) 7.00 d) 4.33  $\therefore$  use  $K_{a1}$  only  
 b) 9.67 e) 3.05  
 c) 7.84  
~~see scratch paper~~
- What is the concentration of  $CO_3^{2-}$  in a 0.010 M solution of carbonic acid? The relevant equilibria are:  
 $H_2CO_3 \rightleftharpoons H^+ + HCO_3^-$   $K_{a1} = 4.3 \times 10^{-7}$   
 $HCO_3^- \rightleftharpoons H^+ + CO_3^{2-}$   $K_{a2} = 5.6 \times 10^{-11}$   
 a)  $6.6 \times 10^{-5} M$  d)  $7.5 \times 10^{-7} M$   
 b)  $5.6 \times 10^{-11} M$  e)  $7.9 \times 10^{-7} M$   
 c)  $6.7 \times 10^{-11} M$   
 From the 1st dissociation,  $[H^+] = [HCO_3^-]$   
 the 2nd dissociation doesn't change this much.  
 $K_{a2} = \frac{[H^+][CO_3^{2-}]}{[HCO_3^-]}$  but  $[H^+] = [HCO_3^-]$  cancel!  
 so...  $K_{a2} = [CO_3^{2-}] = 5.6 \times 10^{-11}$

# Answer Key To Acids Bases And Solutions

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