

Name: Date:

Student Exploration: Stoichiometry

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, conversion factor, dimensional analysis, molar mass, mole, molecular mass, stoichiometry

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? **176g**
2. It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? **15 sticks**

Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The **Stoichiometry** Gizmo allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:



1. Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

Fe_2O_3 : **1** CO : **3** Fe : **2** CO_2 : **3**

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

Reactants	Iron:	2	Oxygen:	6	Carbon:	3
Products	Iron:	2	Oxygen:	6	Carbon:	3

Based on these values, is the equation balanced? **Yes**

New equation:



New question:

How many moles of carbon monoxide (CO) are required to react completely with 1.75 moles of iron (III) oxide (Fe_2O_3)?

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Annelies Wilder-Smith

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