

6**Chemical Quantities**
Downloaded**A. Completion**

They also encourage students to check prior knowledge of the terms and gain understanding of the concepts introduced in this chapter. Each blank can be completed with a unit, three places, or nothing.

Chemists define units of counting, of mass, and of volume as a single quantity called the 1. Chemists measure the "amount of substance" (the weight) measured the amount of a substance by dividing the quantity of 2 of that substance. The amount of representative particles for a sample of a substance is called 3. This number is 4. For example, a mole of nitrogen contains $6.02 \times 10^{23} 5 of nitrogen.$

Chemists define gram atomic mass as the weight of grams of an element whose atoms equal to the 6 atoms. This is that 12.01 g of carbon contains the same number of atoms as 7 g of oxygen atoms.

To find the amount in moles of a substance represented, we will require the 8 of the above-stating equal analogy. When the substance the substance has mass, we divide the 9 of the compound. If the substance is not an appropriate quantity, because the representative particles found in one compound is the formula unit, the representative particle is called as 10. The formula unit mass is called 11 unit. We used to call it the mass of a mole of any substance or compound.

The gram formula mass is a mass of a substance composed of a molecule or a unit of a substance or a formula unit. The definition that mass of a substance is a compound or a mixture, for example, we can find the formula mass as 12. For example, find all the types, find the value for the number of atoms 13 for the substance.

1 mass2 representative particles3 Avogadro's Number4 6.02×10^{23} 5 molecules6 atomic mass7 168 gram atomic masses9 gram molecular mass10 gram formula mass11 Molar mass12 conversion factor13 chemical formula

Answer Key For Chemical Quantities Test

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