

15.1

ELECTRON CONFIGURATION IN IONIC BONDING

SECTION REVIEW

Objectives

- Use the periodic table to infer the number of valence electrons for an atom and draw its electron dot structure.
- Describe the formation of cations from metals and of anions from nonmetals.

Key Terms

- valence electrons
- electron dot structures
- metal cation
- halide ions

Part A Completion

After their completion, compare to check your understanding of the concepts and terms. Check your list against the list in section. Each blank can be completed with a term, other phrase, or name here.

Elements within the same group of the periodic table tend to share many features. They have the same number of _____ 1. _____. This _____ 2. _____ number of a region's valence electrons indicates how many valence electrons that element has. Gilbert Lewis _____ 3. _____ states that in forming compounds, atoms tend to achieve the electron configuration of a noble gas.

Atoms are held together by _____ 4. _____ by chemical bonds. The transfer of valence electrons produces positively charged ions, or _____ 5. _____, and negatively charged ions called _____ 6. _____. The cations of Group 1A elements always have a charge of _____ 7. _____. _____ 8. _____ are produced when atoms of the elements in Group 1A _____ 9. _____ an electron. For transition metals, the _____ 10. _____ of cations may vary.

- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____

Part B True-False

Classify each of these statements as always true, AC (sometimes true), SE (never true), or F.

- _____ 11. The chlorine atom gains seven electrons when it becomes an ion.
- _____ 12. The chemical properties of an element are largely determined by the number of valence electrons that element has.

Answer Key Electron Configuration In Ionic Bonding

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Smith, Jack Price**



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