

## Calculating Equilibrium Concentrations

Given the Equilibrium Constant AND the initial conditions, equilibrium concentrations can be calculated.

Example:

The equilibrium constant for the reaction



is 0.021 at a certain temperature. If the initial concentration of  $\text{PCl}_5(\text{g})$  is 0.100 mol L<sup>-1</sup>, what are the equilibrium concentrations of all species?

Assume a volume of 1 L.

Initially it contains 0.100 moles  $\text{PCl}_5(\text{g})$ .

At equilibrium,  $x$  moles of  $\text{PCl}_5(\text{g})$  have reacted and therefore  $0.1-x$  moles are left. From the equation  $x$  moles of  $\text{PCl}_3(\text{g})$  and of  $\text{Cl}_2(\text{g})$  have been formed.

## 18 3 Calculating Equilibrium Constants

**Jessica J Manson**

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