

- As a chemical bond forms between two hydrogen atoms in a system, energy is released and the stability of the system
  - decreases
  - increases**
  - remains the same
- As two atoms of hydrogen combine to form a molecule of hydrogen, the total energy of the two atoms
  - decreases
  - increases
  - remains the same**
- What is conserved during a chemical reaction?
  - mass, only
  - charge, only
  - both mass and charge**
  - neither mass nor charge
- As a chemical bond forms between two hydrogen atoms, the potential energy of the atoms
  - decreases**
  - increases
  - remains the same
- As energy is released during the formation of a bond, the stability of the chemical system generally will
  - decrease
  - increase**
  - remain the same
- When a sodium atom reacts with a chlorine atom to form a compound, the electron configurations of the ions forming the compound are the same as those in which noble gas atoms?
  - krypton and neon
  - krypton and argon
  - neon and helium
  - neon and argon**
- Which electron-dot diagram represents  $H_2$ ?
  - $H \cdot H$
  - $H : H$
  - $\begin{array}{cc} \cdot & \cdot \\ : & : \\ H & \cdot H \\ \cdot & \cdot \end{array}$
  - $\begin{array}{cc} \cdot & \cdot & \cdot & \cdot \\ : & : & : & : \\ H & : & : & H \\ \cdot & \cdot & \cdot & \cdot \end{array}$
- Which is the correct electron-dot formula for a molecule of chlorine?
  - $\begin{array}{cc} \cdot & \cdot \\ \cdot Cl & : Cl \cdot \\ \cdot & \cdot \end{array}$
  - $\begin{array}{cc} \cdot & \cdot \\ : Cl & : : Cl : \end{array}$
  - $\begin{array}{cc} \cdot & \cdot \\ : Cl & : : Cl : \\ \cdot & \cdot \end{array}$
  - $\begin{array}{cc} \cdot & \cdot \\ : Cl & : Cl : \\ \cdot & \cdot \end{array}$
- A barium atom attains a stable electron configuration when it bonds with
  - one chlorine atom
  - two chlorine atoms**
  - one sodium atom
  - two sodium atoms
- Which pair of elements below will form a compound with the greatest ionic character?
  - Pb and F
  - Ca and O**
  - Na and Cl
  - Cs and N
- Which pair of elements forms a bond with the *least* ionic character?
  - P-Cl
  - Br-Cl**
  - H-Cl
  - O-Cl
- Based on your Reference Tables, the atoms of which of these elements have the strongest attraction for electrons in a chemical bond?
  - N
  - Na
  - P
  - Pt
- As a chlorine atom becomes a negative ion, the atom
  - gains an electron and its radius increases**
  - gains an electron and its radius decreases
  - loses an electron and its radius increases
  - loses an electron and its radius decreases
- Which compound contains ionic bonds?
  - NO
  - NO<sub>2</sub>
  - CaO**
  - CO<sub>2</sub>
- The bonds in BaO are best described as
  - covalent, because valence electrons are shared
  - covalent, because valence electrons are transferred
  - ionic, because valence electrons are shared
  - ionic, because valence electrons are transferred**
- A substance that has a melting point of 1074 K, conducts electricity when dissolved in water, but does *not* conduct electricity in the solid phase. The substance is most likely
  - an ionic solid**
  - a network solid
  - a metallic solid
  - a molecular solid
- Which substance is an electrolyte?
  - CH<sub>3</sub>OH
  - C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>
  - H<sub>2</sub>O
  - KOH**

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