

1. As a chemical bond forms between two hydrogen atoms in a system, energy is released and the stability of the system

A) decreases B) increases
C) remains the same

2. As two atoms of hydrogen combine to form a molecule of hydrogen, the total energy of the two atoms

A) decreases B) increases
C) remains the same

3. What is conserved during a chemical reaction?

A) mass, only
B) charge, only
C) both mass and charge
D) neither mass nor charge

4. As a chemical bond forms between two hydrogen atoms, the potential energy of the atoms

A) decreases B) increases
C) remains the same

5. As energy is released during the formation of a bond, the stability of the chemical system generally will

A) decrease B) increase
C) remain the same

6. When a sodium atom reacts with a chlorine atom to form a compound, the electron configurations of the ions forming the compound are the same as those in which noble gas atoms?

A) krypton and neon B) krypton and argon
C) neon and helium D) neon and argon

7. Which electron-dot diagram represents H_2 ?

A) $\text{H} \cdot \text{H}$ B) $\text{H} : \text{H}$
C) $\begin{array}{c} \cdot \text{H} \cdot \\ \cdot \text{H} \cdot \end{array}$ D) $\begin{array}{c} \cdot \text{H} : \\ \cdot \text{H} : \end{array}$

8. Which is the correct electron-dot formula for a molecule of chlorine?

A) $\begin{array}{c} \cdot \cdot \\ \cdot \text{Cl} : \text{Cl} : \\ \cdot \cdot \end{array}$ B) $\begin{array}{c} \cdot \cdot \cdot \\ \cdot \text{Cl} : \text{Cl} : \\ \cdot \cdot \end{array}$
C) $\begin{array}{c} \cdot \cdot \cdot \\ \cdot \text{Cl} : \text{Cl} : \\ \cdot \cdot \cdot \end{array}$ D) $\begin{array}{c} \cdot \cdot \cdot \\ \cdot \text{Cl} : \text{Cl} : \\ \cdot \cdot \cdot \end{array}$

9. A barium atom attains a stable electron configuration when it bonds with

A) one chlorine atom B) two chlorine atoms
C) one sodium atom D) two sodium atoms

10. Which pair of elements below will form a compound with the greatest ionic character?

A) Pb and F B) Ca and O
C) Na and Cl D) Cs and N

11. Which pair of elements forms a bond with the *least* ionic character?

A) P-Cl B) Br-Cl
C) H-Cl D) O-Cl

12. Based on your Reference Tables, the atoms of which of these elements have the strongest attraction for electrons in a chemical bond?

A) N B) Na C) P D) Pt

13. As a chlorine atom becomes a negative ion, the atom

A) gains an electron and its radius increases
B) gains an electron and its radius decreases
C) loses an electron and its radius increases
D) loses an electron and its radius decreases

14. Which compound contains ionic bonds?

A) NO B) NO_2 C) CaO D) CO_2

15. The bonds in BaO are best described as

A) covalent, because valence electrons are shared
B) covalent, because valence electrons are transferred
C) ionic, because valence electrons are shared
D) ionic, because valence electrons are transferred

16. A substance that has a melting point of 1074 K, conducts electricity when dissolved in water, but does *not* conduct electricity in the solid phase. The substance is most likely

A) an ionic solid B) a network solid
C) a metallic solid D) a molecular solid

17. Which substance is an electrolyte?

A) CH_3OH B) $\text{C}_6\text{H}_{12}\text{O}_6$
C) H_2O D) KOH

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