

1. A 33.7-g sample of HF is dissolved in water to give 2.0×10^2 mL of solution. The concentration of the solution is:

A) 1.68 M
B) 0.34 M
C) 0.17 M
D) 8.4 M
E) 16.9 M

$$[HF] = \frac{33.7 \text{ g HF}}{2.0 \times 10^2 \text{ mL}} \times \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{1 \text{ mol HF}}{20.06 \text{ g HF}} =$$

$$[HF] = 8.4 \text{ M}$$

2. What mass of calcium chloride, CaCl_2 , is needed to prepare 3.419 L of a 1.56 M solution?

A) 243 g
B) 5.33 g
C) 50.6 g
D) 111 g
E) 592 g

$$\frac{\text{mole CaCl}_2}{3.419 \text{ L}} \times \frac{1.56 \text{ mol CaCl}_2}{1 \text{ L}} \times \frac{110.98 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 592 \text{ g CaCl}_2$$

3. A 51.24-g sample of $\text{Ba}(\text{OH})_2$ is dissolved in enough water to make 1.205 liters of solution. How many mL of this solution must be diluted with water in order to make 1.000 liter of 0.1000 molar $\text{Ba}(\text{OH})_2$?

A) 403 mL
B) 248 mL
C) 24.8 mL
D) 83.0 mL
E) 331 mL

$$[\text{Ba}(\text{OH})_2] = \frac{51.24 \text{ g Ba}(\text{OH})_2}{1.205 \text{ L}} \times \frac{1 \text{ mol Ba}(\text{OH})_2}{171.36 \text{ g Ba}(\text{OH})_2}$$

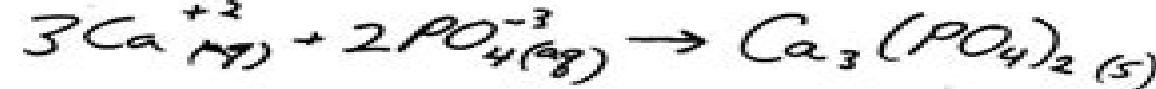
$$[\text{Ba}(\text{OH})_2] = 0.2481 \text{ M}$$

$$M_1 = 0.1000 \text{ M} \quad M_2 = 0.2481 \text{ M} \quad V_1 = \frac{0.1000 \text{ M} \times 1.000 \text{ L}}{0.2481 \text{ M}}$$

$$V_1 = 1.000 \text{ L} \quad V_2 = ? \quad V_2 = 0.4031 \text{ L} = 403 \text{ mL}$$

4. The net ionic equation for the reaction of calcium bromide and sodium phosphate contains which of the following species?

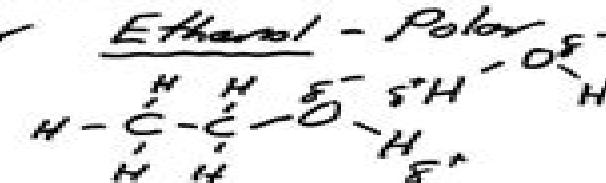
A) $\text{Ca}^{2+}(\text{aq})$
B) $\text{PO}_4^{3-}(\text{aq})$
C) $2\text{Ca}_3(\text{PO}_4)_2(\text{s})$
D) $6\text{NaBr}(\text{aq})$
E) $3\text{Ca}^{2+}(\text{aq})$



5. Consider two organic molecules, ethanol and benzene. One dissolves in water and the other does not. Why?

A) They have different molar masses.
B) One is ionic, the other is not.
C) One is an electrolyte, the other is not.
D) Ethanol contains a polar O—H bond, and benzene does not.
E) Two of these.

*Know the general solubility rule of "Likes Dissolve Likes"



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DOE Budget for Fiscal Year 2011 United States. Congress. Senate. Committee on Energy and Natural Resources, 2010

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Sustainable Municipal Waste Management with a Zero Waste Approach Alok Prasad Das, Kosala Sirisena, Dipalee Bhanja, 2026-01-01 This book includes a variety of cases detailing novel municipal waste management strategies, including the use of microorganisms, harnessing artificial intelligence and machine learning technologies, and implementing advanced waste management techniques. Sustainable municipal waste management, coupled with a zero waste approach and circular economy principles, represents a pivotal strategy for addressing the escalating challenges posed by municipal waste pollution. Municipal waste encompasses a diverse array of materials generated from households, businesses, institutions, and public spaces, including organic waste, plastics, paper, metals, and glass. As urbanization accelerates and consumption patterns evolve, the volume of municipal waste continues to surge, exerting significant pressures on ecosystems, public health, and natural resources. Municipal waste pollution arises from improper disposal practices, inadequate waste management infrastructure, and unsustainable consumption habits. Sources of municipal waste in the environment include landfills, incineration facilities, illegal dumping sites, and littering, contributing to soil, water,

and air pollution The indiscriminate disposal of waste leads to a number of adverse effects and impacts including habitat destruction wildlife entanglement groundwater contamination greenhouse gas emissions and the proliferation of disease vectors To address these challenges innovative approaches blended with sustainability principles are essential With these strategies communities can transition toward more resilient and resource efficient waste management systems **The Physical Chemistry of Electrolytic Solutions** Herbert Spencer Harned,Benton Brooks Owen,1943 **Solutions to Environmental Problems Involving Nanotechnology and Enzyme Technology** Alka Dwevedi,2018-12-06 Nanotechnology and Enzyme Technology Combined to Address Environmental Problems discusses how nanotechnology and enzyme technology work independently and together to help researchers and environmental professionals learn about this revolutionary and cross disciplinary field Nanotechnology has provided a range of nanomaterials some of which are helpful in the protection of the environment and climate They can be used to improve durability against mechanical stress help in cleaning enhance energy efficiency as insulation save energy consumption during transportation due to catalytic properties and more This book highlights this technology as it continues to provide solutions for various environmental problems Covers air and water pollution remediation in the developing field of combining nanotechnology with enzyme technology Reviews the sustainability potentials of combining nanotechnology and enzyme technology including energy production Applies current research and utilization to a variety of environmental issues including pollution and energy production

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