

AP CHEMISTRY MULTIPLE CHOICE QUESTIONS

Unit 8 – Introduction to Equilibrium

KEY



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



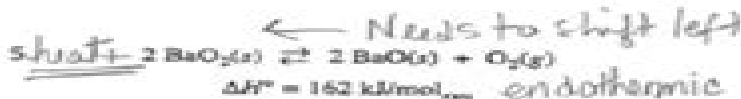
- (A) 95²
 (B) 95
 (C) (95)²
 (D) $\frac{1}{95}$
 (E) $\frac{1}{(95)^2}$



$C_2H_2(g)$ is reduced by $H_2(g)$ in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of $H_2(g)$
 II. Changes in the particle size of the platinum catalyst
 III. Changes in the temperature of the reaction system

- (A) III only
 (B) I and II only
 (C) I and III only
 (D) II and III only
 (E) I, II, and III



A sealed rigid vessel contains $BaO_2(s)$ in equilibrium with $BaO(s)$ and $O_2(g)$ as represented by the equation above. Which of the following changes will increase the amount of $BaO_2(s)$ in the vessel?

- (A) Removing a small amount of $O_2(g)$ *R*
 (B) Removing a small amount of $BaO(s)$ *X*
 (C) Adding He gas to the vessel *X*
 (D) Lowering the temperature *L*

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products? *increase P⁺*

- (A) $SrCO_3(s) \rightleftharpoons SrO(s) + CO_2(g)$
 (B) $Cl_2 + COCl_2(g) \rightleftharpoons 2 COCl(g)$
 (C) $CO(g) + H_2O(g) \rightleftharpoons CO_2(g) + H_2(g)$
 (D) $PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$
 (E) $2 NO_2(g) \rightleftharpoons N_2O_4(g)$
- shift towards fewer & make gas*



When 4.00 mol each of $X(g)$ and $Y(g)$ are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of $Z(g)$ is produced. What is the value of the equilibrium constant, K_c ?

- (A) 3
 (B) 6
 (C) 8
 (D) 16
 (E) 36
- $X + Y \rightleftharpoons 2Z$
 $4M \quad 4M \quad 0$
 $-x \quad -x \quad +2x$
 $1M \quad 1M \quad 6M$
 $2x = 6M \Rightarrow x = 3M$
 $K_c = \frac{[Z]^2}{[X][Y]} = \frac{6^2}{1 \times 1} = 36$

6. The synthesis of $CH_3OH(l)$ from $CO(g)$ and $H_2(g)$ is represented by the equation above. The value of K_c for the reaction at 483 K is 14.5.

Which of the following explains the effect on the equilibrium constant, K_c , when the temperature of the reaction system is increased to 650 K?

- (A) K_c will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.
 (B) K_c will increase because there are more reactant molecules than product molecules.
 (C) K_c will decrease because the reaction is exothermic.
 (D) K_c is constant and will not change.

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