

Version A**AP* Chemistry: 2002 Released Multiple Choice Exam****NO CALCULATORS MAY BE USED**

Note: For all questions, assume that the temperature is 298 K, the pressure is 1.00 atmosphere, and solutions are aqueous unless otherwise specified.

Throughout the test the following symbols have the definitions specified unless otherwise noted.

T	= temperature	M	= molar
P	= pressure	mo	= molar
V	= volume	L , ml	= liter(s), milliliter(s)
S	= entropy	g	= gram(s)
H	= enthalpy	nm	= nanometer(s)
G	= free energy	atm	= atmosphere(s)
R	= molar gas constant	J , kJ	= joule(s), kilojoule(s)
n	= number of moles	V	= volt(s)
mol	= mole(s)		

Directions: Each set of lettered choices below refers to the numbered questions or statements immediately following it. Select the one lettered choice that best answers each question or best fits each statement and then fill in the corresponding oval on the answer sheet. A choice may be used once, more than once, or not at all in each set. *Before turning in your answer sheet, count the number of questions that you have skipped and place that number next to your name ON YOUR ANSWER SHEET and circle it.*

Questions 1-2

Consider atoms of the following elements. Assume that the atoms are in the ground state.

(A) S
(B) Ca
(C) Ga
(D) Sb
(E) Br

1. The atom that contains exactly two unpaired electrons
2. The atom that contains only one electron in the highest occupied energy sublevel

Questions 3-5 refer to the following molecules.

(A) CO_2
(B) H_2O
(C) CH_4
(D) C_2H_4
(E) PH_3

3. The molecule with only one double bond
4. The molecule with the largest dipole moment
5. The molecule that has trigonal pyramidal geometry

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Peter E. Demmin

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