

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa, given K_a or K_b .
2. Calculate K_a and K_b for weak acids and bases, and vice versa, given K_a or K_b .
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|----------------|----------|----------|
| acid | 1. K_a | 2. K_b |
| base | 3. K_a | 4. K_b |
| conjugate acid | 5. K_a | 6. K_b |
| conjugate base | 7. K_a | 8. K_b |

Key Equations

- | | |
|------------------------------------|-------------------------------------|
| 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ | 2. $K_b = \frac{[OH^-][HB^+]}{[B]}$ |
| 3. $K_a = 10^{-pH}$ | 4. $K_b = 10^{-pOH}$ |

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, showing double bonds where necessary. Write the chemical formula for each of the following acids, bases, and salts.

1. $K_a = 1.0 \times 10^{-4}$ (weak acid)
2. $K_b = 1.0 \times 10^{-4}$ (weak base)
3. $K_a = 1.0 \times 10^{-4}$ (weak acid)
4. $K_b = 1.0 \times 10^{-4}$ (weak base)
5. $K_a = 1.0 \times 10^{-4}$ (weak acid)
6. $K_b = 1.0 \times 10^{-4}$ (weak base)
7. $K_a = 1.0 \times 10^{-4}$ (weak acid)
8. $K_b = 1.0 \times 10^{-4}$ (weak base)
9. $K_a = 1.0 \times 10^{-4}$ (weak acid)
10. $K_b = 1.0 \times 10^{-4}$ (weak base)
11. $K_a = 1.0 \times 10^{-4}$ (weak acid)
12. $K_b = 1.0 \times 10^{-4}$ (weak base)
13. $K_a = 1.0 \times 10^{-4}$ (weak acid)
14. $K_b = 1.0 \times 10^{-4}$ (weak base)
15. $K_a = 1.0 \times 10^{-4}$ (weak acid)
16. $K_b = 1.0 \times 10^{-4}$ (weak base)
17. $K_a = 1.0 \times 10^{-4}$ (weak acid)
18. $K_b = 1.0 \times 10^{-4}$ (weak base)
19. $K_a = 1.0 \times 10^{-4}$ (weak acid)
20. $K_b = 1.0 \times 10^{-4}$ (weak base)

1. $K_a = 1.0 \times 10^{-4}$
2. $K_b = 1.0 \times 10^{-4}$
3. $K_a = 1.0 \times 10^{-4}$
4. $K_b = 1.0 \times 10^{-4}$
5. $K_a = 1.0 \times 10^{-4}$
6. $K_b = 1.0 \times 10^{-4}$
7. $K_a = 1.0 \times 10^{-4}$
8. $K_b = 1.0 \times 10^{-4}$
9. $K_a = 1.0 \times 10^{-4}$
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11. $K_a = 1.0 \times 10^{-4}$
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15. $K_a = 1.0 \times 10^{-4}$
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17. $K_a = 1.0 \times 10^{-4}$
18. $K_b = 1.0 \times 10^{-4}$
19. $K_a = 1.0 \times 10^{-4}$
20. $K_b = 1.0 \times 10^{-4}$

19 2 Hydrogen Ions And Acidity Section Review

Linda Swisher, Gary A. Thibodeau



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