

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on  $K_a$  and  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                   |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution  |
| 4. autoionization | 5. weak acid                | 6. conjugate base |
|                   | 7. weak solution            | 8. pH             |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the unknown quantity that are indicated in the spaces provided. Be sure to include correct units when appropriate.

1. What substance is \_\_\_\_\_, a weak hydrogen ion donor?
2. \_\_\_\_\_ is the concentration of those species present in water that equal to \_\_\_\_\_.
3. \_\_\_\_\_ is the concentration of those species present in \_\_\_\_\_.
4. \_\_\_\_\_ is the concentration of those species present in \_\_\_\_\_.
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19. \_\_\_\_\_ is the concentration of those species present in \_\_\_\_\_.
20. \_\_\_\_\_ is the concentration of those species present in \_\_\_\_\_.

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# 19 2 Hydrogen Ions And Acidity Answer Key

**Didier Musso**



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