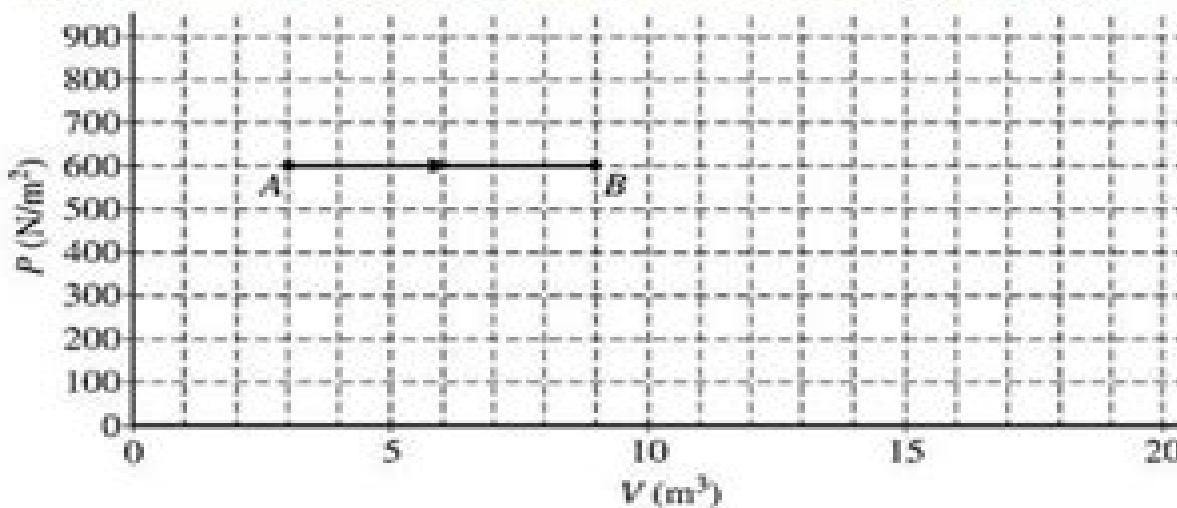


2004 AP[®] PHYSICS B FREE-RESPONSE QUESTIONS



5. (10 points)

The diagram above of pressure P versus volume V shows the expansion of 2.0 moles of a monatomic ideal gas from state A to state B . As shown in the diagram, $P_A = P_B = 600$ N/m², $V_A = 3.0$ m³, and $V_B = 9.0$ m³.

(a)

- Calculate the work done by the gas as it expands.
- Calculate the change in internal energy of the gas as it expands.
- Calculate the heat added to or removed from the gas during this expansion.

(b) The pressure is then reduced to 200 N/m² without changing the volume as the gas is taken from state B to state C . Label state C on the diagram and draw a line or curve to represent the process from state B to state C .

(c) The gas is then compressed isothermally back to state A .

- Draw a line or curve on the diagram to represent this process.
- Is heat added to or removed from the gas during this isothermal compression?

_____ added to _____ removed from

Justify your answer.

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Table of Contents 2000 Ap Physics B Response Answers

1. Understanding the eBook 2000 Ap Physics B Response Answers
 - The Rise of Digital Reading 2000 Ap Physics B Response Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 2000 Ap Physics B Response Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2000 Ap Physics B Response Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2000 Ap Physics B Response Answers
 - Personalized Recommendations
 - 2000 Ap Physics B Response Answers User Reviews and Ratings
 - 2000 Ap Physics B Response Answers and Bestseller Lists
5. Accessing 2000 Ap Physics B Response Answers Free and Paid eBooks
 - 2000 Ap Physics B Response Answers Public Domain eBooks
 - 2000 Ap Physics B Response Answers eBook Subscription Services
 - 2000 Ap Physics B Response Answers Budget-Friendly Options

6. Navigating 2000 Ap Physics B Response Answers eBook Formats
 - ePUB, PDF, MOBI, and More
 - 2000 Ap Physics B Response Answers Compatibility with Devices
 - 2000 Ap Physics B Response Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2000 Ap Physics B Response Answers
 - Highlighting and Note-Taking 2000 Ap Physics B Response Answers
 - Interactive Elements 2000 Ap Physics B Response Answers
8. Staying Engaged with 2000 Ap Physics B Response Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2000 Ap Physics B Response Answers
9. Balancing eBooks and Physical Books 2000 Ap Physics B Response Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2000 Ap Physics B Response Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 2000 Ap Physics B Response Answers
 - Setting Reading Goals 2000 Ap Physics B Response Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2000 Ap Physics B Response Answers
 - Fact-Checking eBook Content of 2000 Ap Physics B Response Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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