

ANSWER KEY: Ch. 17 Acid-base and Solubility Equilibria

Question 1 Which of the following acids is not a weak acid? *1. Which of the following acids is not a weak acid?*

Answer 1 *1. Which of the following acids is not a weak acid?*

The strong, nonvolatile acids and weak acids are listed in concordance with acid-base literature. Which one of the following statements is not true?

- ☐ A. The equivalence point and end point will always be exactly the same for any titration.
- ☐ B. As the concentration of the analyte or titrant in the solution being titrated is the same as the number of equivalents of the titrant added.
- ☐ C. The end point of a titration is often dependent on an indicator color change.
- ☐ D. The equivalence point of a titration can be found using a pH meter.

Question 2

2. How many milliliters of 0.10 M NaOH are needed to neutralize 10 grams of HCl dissolved in water? (HCl is a strong acid.)

Answer 2 *2. How many milliliters of 0.10 M NaOH are needed to neutralize 10 grams of HCl dissolved in water? (HCl is a strong acid.)*

- ☐ A. 1000 mL
- ☐ B. 100 mL
- ☐ C. 250 mL
- ☐ D. 25 mL

Question 3 *3. Which of the following is a weak acid?*

Answer 3 *3. Which of the following is a weak acid?*

- ☐ A. $1.0 \times 10^{-2} \text{ mol/L}$
- ☐ B. $1.0 \times 10^{-2} \text{ mol/L}$

17 1 Answer Key Solubility Equilibria

Andrea Leeson



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