

Colligative Properties Worksheet – Answers

- 1) If I add 45 grams of sodium chloride to 500 grams of water, what will the melting and boiling points be of the resulting solution? $K_b(\text{H}_2\text{O}) = 0.52^\circ\text{C}/\text{m}$ and $K_f(\text{H}_2\text{O}) = 1.86^\circ\text{C}/\text{m}$.

To find the melting and boiling points, you'll first need to find the concentration of the solution. To find the molality, convert grams of sodium chloride to moles (it turns out to be 0.769 moles) and divide by the number of kilograms of solvent (0.500 kg H_2O). When you do this calculation, the molality is 1.54 m. Because sodium chloride breaks into two particles when it dissociates in water, the effective molality for the purposes of colligative property calculations is twice this, or 3.08 m.

From here, it's a matter of plugging the molality into the equations for freezing point depression and boiling point elevation. When you do this, you should get that the change in melting point is 5.73°C and the change in boiling point is 1.60°C .

The final answers: Melting point = -5.73°C , boiling point = 101.60°C .

- 2) What is the vapor pressure of the solution in problem #1 at 25°C ? The vapor pressure of pure water at 25°C is 3.17 kPa.

Vapor pressure = mole fraction of water \times normal vapor pressure of water. The mole fraction of water in this case is equal to the number of moles of water (27.8) divided by the total number of moles of all substances in the solution ($27.8 + 0.769$) = 28.6 moles. The mole fraction of water is 0.972. When you multiply the mole fraction times the normal vapor pressure, you find that the vapor pressure of the solution is 3.08 kPa.

- 3) Which solution will have a higher boiling point: A solution containing 105 grams of sucrose ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$) in 500 grams of water or a solution containing 35 grams of sodium chloride in 500 grams of water?

Using the same calculations as in question 1, you find that the boiling point of the sucrose solution is 100.32°C and the boiling point of the sodium chloride solution is 101.40°C . Clearly, the sodium chloride solution has a higher boiling point.

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Worksheet Answer Key

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Table of Contents 16 4 Calculations Involving Colligative Properties Worksheet Answer Key

1. Understanding the eBook 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
 - The Rise of Digital Reading 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
 - User-Friendly Interface

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4. Exploring eBook Recommendations from 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
 - Personalized Recommendations
 - 16 4 Calculations Involving Colligative Properties Worksheet Answer Key User Reviews and Ratings
 - 16 4 Calculations Involving Colligative Properties Worksheet Answer Key and Bestseller Lists
5. Accessing 16 4 Calculations Involving Colligative Properties Worksheet Answer Key Free and Paid eBooks
 - 16 4 Calculations Involving Colligative Properties Worksheet Answer Key Public Domain eBooks
 - 16 4 Calculations Involving Colligative Properties Worksheet Answer Key eBook Subscription Services
 - 16 4 Calculations Involving Colligative Properties Worksheet Answer Key Budget-Friendly Options
6. Navigating 16 4 Calculations Involving Colligative Properties Worksheet Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - 16 4 Calculations Involving Colligative Properties Worksheet Answer Key Compatibility with Devices
 - 16 4 Calculations Involving Colligative Properties Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
 - Highlighting and Note-Taking 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
 - Interactive Elements 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
8. Staying Engaged with 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
9. Balancing eBooks and Physical Books 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
 - Setting Reading Goals 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
 - Carving Out Dedicated Reading Time

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12. Sourcing Reliable Information of 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
 - Fact-Checking eBook Content of 16 4 Calculations Involving Colligative Properties Worksheet Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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