

(ix) The standard potential of the cell in the following reaction is _____.
 $\text{Cd}_{(\text{ss})} + \text{Cu}_{(\text{ss})}^{2+} \longrightarrow \text{Cd}_{(\text{ss})}^{2+} + \text{Cu}_{(\text{ss})}$
 $(E_{\text{cell}}^{\circ} = -0.403 \text{ V}, E_{\text{cell}}^{\circ} = 0.334 \text{ V})$

(a) -0.737 V (b) 0.737 V
 (c) -0.069 V (d) 0.069 V

(x) The value of $[\text{H}_3\text{O}^+]$ in mol lit^{-1} of 0.001 M acetic acid solution ($K_a = 1.8 \times 10^{-5}$) is _____.
 (a) 1.34×10^{-1} (b) 1.34×10^{-2}
 (c) 1.34×10^{-3} (d) 1.34×10^{-4}

Q.2. Answer the following questions:

[8]

(i) Write the product formed when alkyl halide reacts with silver nitrite.

(ii) Write the name of product formed, when acetone is treated with 2, 4-dinitrophenyl hydrazine.

(iii) Write the name of biodegradable polyamide copolymer.

(iv) Identify the molecularity of following elementary reaction:
 $\text{NO}_{(\text{g})} + \text{O}_{(\text{g})} \longrightarrow \text{NO}_{(\text{g})} + \text{O}_{(\text{g})}$

(v) What is the action of selenium on magnesium metal?

(vi) Write the name of isomerism in the following complexes:
 $[\text{Cu}(\text{NH}_3)_4] [\text{PtCl}_4]$ and $[\text{Pt}(\text{NH}_3)_4] [\text{CuCl}_4]$

(vii) Write the name of the alloy used in Fischer Tropsch process in the synthesis of gasoline.

(viii) Henry's law constant for $\text{CH}_3\text{Br}_{(\text{g})}$ is $0.159 \text{ mol dm}^{-3} \text{ bar}^{-1}$ at 25°C . What is solubility of $\text{CH}_3\text{Br}_{(\text{g})}$ in water at same temperature and partial pressure of 0.164 bar?

SECTION – B

Attempt any EIGHT of the following questions:

[16]

Q.3. Explain pseudo-first order reaction with suitable example.

Q.4. Write the consequences of Schottky defect with reasons.

Q.5. What is the action of following on ethyl bromide:
 (i) Na in dry ether
 (ii) Mg in dry ether

Q.6. Explain formation of peptide linkage in protein with an example.

Q.7. Derive an expression to calculate molar mass of non volatile solute by osmotic pressure measurement.

Q.8. Explain monodentate and ambidentate ligands with example.

Q.9. Explain the trends in the following atomic properties of group 16 elements:

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