

(ix) The standard potential of the cell in the following reaction is _____



($E_{\text{Cd}^{2+}}^{\circ} = -0.403 \text{ V}$, $E_{\text{Cu}^{2+}}^{\circ} = 0.334 \text{ V}$)

(a) -0.737 V

(b) 0.737 V

(c) -0.069 V

(d) 0.069 V

(x) The value of $[\text{H}_3\text{O}^+]$ in mol lit^{-1} of 0.001 M acetic acid solution ($K_a = 1.8 \times 10^{-5}$) is _____

(a) 1.34×10^{-1}

(b) 1.34×10^{-2}

(c) 1.34×10^{-3}

(d) 1.34×10^{-4}

Q.2. Answer the following questions:

[8]

(i) Write the product formed when alkyl halide reacts with silver nitrite.

(ii) Write the name of product formed, when acetone is treated with 2, 4-dinitrophenyl hydrazine.

(iii) Write the name of biodegradable polyamide copolymer.

(iv) Identify the molecularity of following elementary reaction:



(v) What is the action of selenium on magnesium metal?

(vi) Write the name of isomerism in the following complexes:



(vii) Write the name of the alloy used in Fischer Tropsch process in the synthesis of gasoline.

(viii) Henry's law constant for $\text{CH}_3\text{Br}_{(g)}$ is $0.159 \text{ mol dm}^{-3} \text{ bar}^{-1}$ at 25°C . What is solubility of $\text{CH}_3\text{Br}_{(g)}$ in water at same temperature and partial pressure of 0.164 bar ?

SECTION – B

Attempt any EIGHT of the following questions:

[16]

Q.3. Explain pseudo-first order reaction with suitable example.

Q.4. Write the consequences of Schottky defect with reasons.

Q.5. What is the action of following on ethyl bromide:

(i) Na in dry ether

(ii) Mg in dry ether

Q.6. Explain formation of peptide linkage in protein with an example.

Q.7. Derive an expression to calculate molar mass of non volatile solute by osmotic pressure measurement.

Q.8. Explain monodentate and ambidentate ligands with example.

Q.9. Explain the trends in the following atomic properties of group 16 elements:

(i) Atomic radii

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Table of Contents 12th Science Question Paper

1. Understanding the eBook 12th Science Question Paper
 - The Rise of Digital Reading 12th Science Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying 12th Science Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 12th Science Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from 12th Science Question Paper
 - Personalized Recommendations
 - 12th Science Question Paper User Reviews and Ratings
 - 12th Science Question Paper and Bestseller Lists

5. Accessing 12th Science Question Paper Free and Paid eBooks
 - 12th Science Question Paper Public Domain eBooks
 - 12th Science Question Paper eBook Subscription Services
 - 12th Science Question Paper Budget-Friendly Options
6. Navigating 12th Science Question Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - 12th Science Question Paper Compatibility with Devices
 - 12th Science Question Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 12th Science Question Paper
 - Highlighting and Note-Taking 12th Science Question Paper
 - Interactive Elements 12th Science Question Paper
8. Staying Engaged with 12th Science Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 12th Science Question Paper
9. Balancing eBooks and Physical Books 12th Science Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 12th Science Question Paper
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 12th Science Question Paper
 - Setting Reading Goals 12th Science Question Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 12th Science Question Paper
 - Fact-Checking eBook Content of 12th Science Question Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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