

7.10 Compounds of sulfur :

7.10.1 Sulfur dioxide

a. Preparation :

i. **From sulfur** : Sulfur dioxide gas can be prepared by burning of sulfur in air.



ii. **From sulfite** : In the laboratory sulfur dioxide is prepared by treating aqueous solution of sodium sulfite with dilute sulfuric acid.



iii. From sulfides : (Industrial method)

Sulfur dioxide can be prepared by roasting zinc sulfide and iron pyrites.



b. Physical properties of SO_2

i. Sulfur dioxide is a colourless gas with a pungent smell.

ii. It is poisonous in nature.

iii. SO_2 is highly soluble in water and its solution in water is called sulfurous acid.

iv. It liquifies at room temperature under a pressure of 2 atm and boils at 263 K.

c. Chemical Properties :

i. **Reaction with Cl_2** : Sulfur dioxide reacts with chlorine in the presence of charcoal (catalyst) to form sulfuryl chloride.



ii. **Reaction with O_2** : Sulfur dioxide is oxidised by dioxygen in presence of vanadium (V) oxide to sulfur trioxide.



iii. **Reaction with NaOH** : Sulfur dioxide readily reacts with sodium hydroxide solution to form sodium sulfite.



iv. **Reaction with Na_2SO_3** : Sulfur dioxide reacts with sodium sulfite solution to form sodium hydrogen sulfite.



v. **Reducing property** : Sulfur dioxide acts as a reducing agent in the presence of moisture.

- Moist sulfur dioxide reduces ferric salts into ferrous salts.



- Moist sulfur dioxide decolourises acidified potassium permanganate (VII) solution.



- Moist sulfur dioxide reduces halogens to halogen acids.



d. **Structure of SO_2** : Sulfur dioxide is angular with $\text{O} = \text{S} = \text{O}$ bond angle of 119.5° .



The $\text{S} = \text{O}$ double bond arises from $\text{dx} - \text{py}$ bonding. It is a resonance hybrid of two canonical forms.

e. Uses : Sulfur dioxide is used

- In refining of petroleum and sugar.
- In bleaching wool and silk.
- As an anti-chlor, disinfectant.
- As a preservative.
- In the manufacture of H_2SO_4 , NaHSO_3 .
- Liquid SO_2 is used as a solvent to dissolve a number of organic and inorganic chemicals.

7.10.2 Sulfuric acid, H_2SO_4

a. **Preparation** : Sulfuric acid is manufactured by **Contact process**, which involves the following three steps.

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