

Chemistry

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 (Chapter - 3) (Chemical Kinetics) (Exercises)
 (Class 12)

Question 3.1:

From the rate expression for the following reactions, determine their order of reaction and the dimensions of the rate constants.

- (i) $3\text{NO(g)} \rightarrow \text{N}_2\text{O(g)}$ Rate = $k[\text{NO}]^2$
- (ii) $\text{H}_2\text{O}_2\text{(aq)} + 3\text{I}^-(\text{aq}) + 2\text{H}^+ \rightarrow 2\text{H}_2\text{O(l)} + \text{I}_3^-$ Rate = $k[\text{H}_2\text{O}_2][\text{I}^-]^3$
- (iii) $\text{CH}_3\text{CHO(g)} \rightarrow \text{CH}_4\text{(g)} + \text{CO(g)}$ Rate = $k[\text{CH}_3\text{CHO}]^{3/2}$
- (iv) $\text{C}_2\text{H}_5\text{Cl(g)} \rightarrow \text{C}_2\text{H}_4\text{(g)} + \text{HCl(g)}$ Rate = $k[\text{C}_2\text{H}_5\text{Cl}]$

Answer 3.1:

(i) Given rate = $k[\text{NO}]^2$

Therefore, order of the reaction = 2

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{NO}]^2} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{(\text{mol L}^{-1})^2} = \frac{\text{mol L}^{-1}\text{s}^{-1}}{\text{mol}^2\text{L}^{-2}} = \text{L mol}^{-1}\text{s}^{-1}$$

(ii) Given rate = $k[\text{H}_2\text{O}_2][\text{I}^-]$

Therefore, order of the reaction = 2

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{H}_2\text{O}_2][\text{I}^-]} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{(\text{mol L}^{-1})(\text{mol L}^{-1})} = \text{L mol}^{-1}\text{s}^{-1}$$

(iii) Given rate = $k[\text{CH}_3\text{CHO}]^{3/2}$

Therefore, order of reaction = 3/2

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{CH}_3\text{CHO}]^{3/2}} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{(\text{mol L}^{-1})^{3/2}} = \frac{\text{mol L}^{-1}\text{s}^{-1}}{\text{mol}^{3/2}\text{L}^{3/2}} = \text{L}^{3/2}\text{mol}^{1/2}\text{s}^{-1}$$

(iv) Given rate = $k[\text{C}_2\text{H}_5\text{Cl}]$

Therefore, order of the reaction = 1

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{C}_2\text{H}_5\text{Cl}]} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{\text{mol L}^{-1}} = \text{s}^{-1}$$

Question 3.2:

For the reaction:



The rate = $k[\text{A}][\text{B}]^2$ with $k = 2.0 \times 10^{-6}\text{mol}^{-2}\text{L}^2\text{s}^{-1}$. Calculate the initial rate of reaction when

$[\text{A}] = 0.1\text{ mol L}^{-1}$, $[\text{B}] = 0.2\text{ mol L}^{-1}$. Calculate the rate of reaction after $[\text{A}]$ is reduced to 0.06 mol L^{-1}

Answer 3.2:

The initial rate of reaction is

$$\text{Rate} = k[\text{A}][\text{B}]^2 \Rightarrow (2.0 \times 10^{-6}\text{mol}^{-2}\text{L}^2\text{s}^{-1})(0.1\text{ mol L}^{-1})(0.2\text{ mol L}^{-1})^2 = 8.0 \times 10^{-9}\text{mol}^{-2}\text{L}^2\text{s}^{-1}$$

When $[\text{A}]$ is reduced from 0.1 mol L^{-1} to 0.06 mol L^{-1} ,

$$\text{The concentration of A reacted} = (0.1 - 0.06)\text{mol L}^{-1} = 0.04\text{ mol L}^{-1}$$

$$\text{Therefore, concentration of B reacted} = \frac{1}{2} \times 0.04\text{ mol L}^{-1}$$

$$\text{Then concentration of B available, } [\text{B}] = (0.2 - 0.02)\text{mol L}^{-1} = 0.18\text{ mol L}^{-1}$$

After $[\text{A}]$ is reduced to 0.06 mol L^{-1} , the rate of reaction is given by,

$$\text{Rate} = k[\text{A}][\text{B}]^2 \Rightarrow (2.0 \times 10^{-6}\text{mol}^{-2}\text{L}^2\text{s}^{-1})(0.06\text{ mol L}^{-1})(0.18\text{ mol L}^{-1})^2 = 3.89 \text{ mol L}^{-1}\text{s}^{-1}$$

Question 3.3:

The decomposition of NH_3 on platinum is zero order reaction. What are the rates of production of N_2 and H_2 if $k = 2.5 \times 10^{-8}\text{ mol}^{-1}\text{ L s}^{-1}$?

Answer 3.3:

The decomposition of NH_3 on platinum surface is represented by the following equation.



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JL Elias

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