

**CHEMISTRY**  
**Class – XII**

Time – 3 hrs.

M.M. : 70

Date – 14.09.2022 (Wednesday)

Name of the student \_\_\_\_\_

Section \_\_\_\_\_

**GENERAL INSTRUCTIONS:**

- All questions are compulsory.
- Question nos. 1 to 21 are MCQs with one correct option and carrying 1 mark each.
- Question nos. 22 to 24 are objective/short answer type questions.
- Question nos. 25 to 28 are of 2 marks each.
- Question nos. 29 to 32 is of 3 marks each.
- Question nos. 33 to 35 is of 5 marks each.

**SECTION - A**

**Q. 1 to 21 is of 1 mark each.**

**Q.1** Which of the following aqueous solution should have highest boiling point ?

- a) 1 M NaOH                      b) 1 M Na<sub>2</sub>SO<sub>4</sub>                      c) 1M NH<sub>4</sub>NO<sub>3</sub>                      d) 1 M KNO<sub>3</sub>

**Q.2** Which of the following statement is false ?

- a) Two different solution of sucrose of same molality prepare in different solvent will have the same depression in freezing point.
- b) The osmotic pressure of a solution is given by the equation  $\pi = CRT$ . (where C is the molarity of the solution)
- c) Decreasing order of osmotic pressure for 0.01M aqueous solution of Barium Chloride, Potassium chloride, acetic acid and sucrose is BaCl<sub>2</sub> > KCl > CH<sub>3</sub>COOH > Sucrose
- d) According to Raoult's Law, the vapour pressure exerted by a volatile component of solution is directly proportional to its molar fraction in the solution.

**Q.3** We have three aqueous solution of NaCl labelled as 'A', 'B' and 'C' with concentration 0.1 M, 0.01 M and 0.001 M respectively. The value of Van't Hoff factor for these solution will be in the order-

- a)  $i_A < i_B < i_C$                       a)  $i_A > i_B > i_C$                       a)  $i_A = i_B = i_C$                       a)  $i_A < i_B > i_C$

**Q.4** On the basis of the information given below mark the correct options –

- i) In bromoethane and chloroethane mixture intermolecular interaction of A-A and B-B type are nearly same as A-B type interaction.
- ii) In the ethanol and acetone mixture A-A or B-B type intermolecular interactions are stronger than A-B type interaction.

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