

KEY

5650 kJ

1. The combustion of 1.00 mol of sucrose,  $C_{12}H_{22}O_{11}$ , evolves  $6.65 \times 10^3$  kJ of heat. A bomb calorimeter has a calibrated heat capacity of  $1.23 \text{ kJ}^\circ\text{C}$ . How many grams of sucrose should be burned to raise the temperature of the calorimeter and its contents from  $23.0^\circ\text{C}$  to  $80.0^\circ\text{C}$ ?  $\Delta T = 80.0^\circ - 23.0^\circ = 57.0^\circ\text{C}$

$$q = C_{\text{cal}} \cdot \Delta T = 1.23 \frac{\text{kJ}}{\text{C}} \cdot 57.0^\circ\text{C} = 70.11 \text{ kJ \text{ needed}}$$

$$1 \text{ mole sucrose} = 342 \cdot 3 \text{ g} ; \frac{342 \cdot 3 \text{ g}}{5650 \text{ kJ}} = \frac{x \text{ g}}{70.11 \text{ kJ}} ; x = \boxed{4.25 \text{ g sucrose}}$$

\* don't have to do this.

The heat of reaction for burning 1 mole of a certain compound X is known to be  $-477.7$  kJ. The calorimeter constant of the bomb being used is  $2.5 \times 10^3 \text{ J}^\circ\text{C}$  and the initial temperature of the water is  $23.2^\circ\text{C}$ .

(A) If  $22.54$  g of compound X ( $M_M = 46.0$ ) is burned in the bomb calorimeter containing  $2000$  mL of water ( $S.H. = 4.184 \text{ J/g}^\circ\text{C}$ ), what will be the final temperature?

$$T_f = 44.7^\circ\text{C}$$

(B) How much water can be warmed from  $23.2^\circ\text{C}$  to  $56.5^\circ\text{C}$  when  $52.3$  g of the compound is burned in the bomb?

$$13300 \text{ g H}_2\text{O}$$

3. A  $500.0$ -g piece of metal at  $60.0^\circ\text{C}$  is placed in  $200.0$  g of water at  $22.0^\circ\text{C}$  contained in a coffee cup calorimeter. The metal and water come to the same temperature at  $32.5^\circ\text{C}$ . The specific heat of water is  $4.184 \text{ J/g}^\circ\text{C}$

$$q_{\text{lost by metal}} = q_{\text{gained by water}}$$

$$\Delta T_{\text{water}} = 32.5^\circ - 22.0^\circ = 10.5^\circ\text{C}$$

(A) How much heat energy did the metal give up to the water?

$$q_{\text{water}} = (200.0 \text{ g})(4.184 \frac{\text{J}}{\text{g}^\circ\text{C}})(10.5^\circ\text{C}) = \boxed{8786 \text{ J released by metal}}$$

## 12 4 Review And Reinforcement Calorimetry Answers

**Richard B. Kemp**

## 12 4 Review And Reinforcement Calorimetry Answers:

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**Handbook of Thermal Analysis and Calorimetry** Richard B. Kemp, 1999-12-13 The applications and interest in thermal analysis and calorimetry have grown enormously during the last half of the 20th century. These techniques have become indispensable in the study of processes such as catalysis hazards evaluation etc and in measuring important physical properties quickly conveniently and with markedly improved accuracy. Consequently thermal analysis and calorimetry have grown in stature and more scientists and engineers have become at least part time practitioners. People new to the field therefore need a source of information describing the basic principles and current state of the art. The last volume of this 4 volume handbook devoted to many aspects of biological thermal analysis and calorimetry completes a comprehensive review of this important area. All chapters have been prepared by recognized experts in their respective fields. The approach taken is how and what to do and when to do it. The complete work is a valuable addition to the already existing literature

**Analytical Calorimetry** Roger Porter, 2012-12-06 Thermal analysis cuts a broad swathe through contemporary science. Within this domain advances in instrumentation permit the application of quantitative calorimetry to the full spectrum of modern materials. This can be illustrated perhaps no better than by the set of contributions which make up this Volume 4 of Analytical Calorimetry. The research supported in this fourth volume of Analytical Calorimetry covers a wide variety of topics. The range indicates the sophistication which thermal analysis is reaching and additionally the ever widening applications that are being developed. The contributions to the Volume represent in part papers presented before the Division of Analytical Chemistry at the Fourth Symposium on Analytical Chemistry held at the 172nd National Meeting of the American Chemical Society August 29 September 3 1976. Several additional key contributions are included by prominent workers in the field Julian F Johnson Roger S Porter v Quantitative Scanning Calorimetry with Computerized Data Acquisition and Reduction 1 Aris Doelman Annie R Gregges and Edward M Barrall II Thermodynamic Properties of a Series of P N Alkyl P Cyanobiphenyls 19 E M Barrall II R J Cox A Doelman N Cleckak J A Logan and A R Gregges Thermodynamic Properties of a Series of 37 P N Alkoxy P Cyanobiphenyls R J Cox E M Barrall II A Doelman N J Cleckak J A Logan R D Diller and A R Gregges A New Equation for Predicting Polyethylene Lifetimes from Thermal Data 51 Cary F

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principles experimental apparatus and practical procedures used in thermal analysis and calorimetric methods of analysis Brief accounts of the basic theory are reinforced with detailed applications of the methods and contemporary developments Also included is information on standard test methods and manufacturers Written by acknowledged experts Principles of Thermal Analysis and Calorimetry is up to date wide ranging and practical It will be an important source of information for many levels of readership in a variety of areas from students and lecturers through to industrial and laboratory staff and consultants     *Calorimetry* Wojciech Zielenkiewicz,2005

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